

NEW MEXICO STATE LAND OFFICE  
 OFFICE OF THE STATE GEOLOGIST  
 SANTA FE, NEW MEXICO

MISCELLANEOUS REPORTS ON WELLS

Submit this report in duplicate to the State Geologist or proper Oil and Gas Inspector within ten days after the work specified is completed. It should be signed and sworn to before a notary public for reports on beginning drilling operations, results of shooting well, results of test of water shut-off, result of abandonment of well, and other important operations, even though the work was witnessed by the State Geologist or Oil and Gas Inspector. Reports on minor operations need not be signed and sworn to before a notary public, but such operations should be witnessed by an Oil and Gas Inspector if possible.

Indicate nature of report by checking below:

REPORT ON BEGINNING DRILLING OPERATIONS		REPORT ON DEEPENING WELL	
REPORT ON RESULT OF SHOOTING WELL	X	REPORT ON PULLING OR OTHERWISE ALTERING CASING	
REPORT ON RESULT OF TEST OF WATER SHUT-OFF		REPORT ON REPAIRING WELL	
REPORT ON RESULT OF ABANDONMENT OF WELL			

Lea Pool 12/28/34

Mr. E. H. Wells State Geologist,  
 Santa Fe, N. Mex.

PLACE DATE

Following is a report on the work done and the results obtained under the heading noted above at the General Crude Oil Company State B Well No. 4 in the

NW. SW of Sec. 2, T. 21s, R. 35E, N. M. P. M.,  
Lea Pool Oil Field, Lea County.

The dates of this work were as follows:

Notice of intention to do the work was ~~(was not)~~ submitted on Form SG 105 on 12/18/34, 19    , and approval of the proposed plan was (was not) obtained. (Cross out incorrect words.)

DETAILED ACCOUNT OF WORK DONE AND RESULTS OBTAINED

Treated our State B # 1 with 1500 gallons of acid (Cold) on the 12/12/34. set for 72 hours. Result obtained - a increase of 10% in production. P/L. 0.

DUPLICATE

Subscribed and sworn to before me this \_\_\_\_\_ day of \_\_\_\_\_, 19\_\_\_\_\_.

I hereby swear or affirm that the information given above is true and correct.

Name M. J. Gray  
 Position Dist. Supt.  
 Representing General Crude Oil Co.

NOTARY PUBLIC.

My commission expires \_\_\_\_\_

Address Wink, Texas. Box 665.

Remarks:

N.C.R.

APPROVED AS TRUE BY [Signature]

1-8-35  
 APPROVED AS CORRECT BY [Signature]

TITLE

THE UNIVERSITY OF CHICAGO  
DEPARTMENT OF CHEMISTRY  
RESEARCH REPORT

The following is a summary of the work done in the laboratory of Professor [Name] during the year 1954. The work was supported by the National Science Foundation and the University of Chicago.

The first part of the work was devoted to the study of the reaction of [Chemical] with [Chemical]. It was found that the reaction proceeds through a series of steps, the first of which is the formation of a complex. The rate of reaction is first order in [Chemical] and second order in [Chemical]. The activation energy of the reaction is [Value] kcal/mole.

The second part of the work was devoted to the study of the reaction of [Chemical] with [Chemical]. It was found that the reaction proceeds through a series of steps, the first of which is the formation of a complex. The rate of reaction is first order in [Chemical] and second order in [Chemical]. The activation energy of the reaction is [Value] kcal/mole.

The third part of the work was devoted to the study of the reaction of [Chemical] with [Chemical]. It was found that the reaction proceeds through a series of steps, the first of which is the formation of a complex. The rate of reaction is first order in [Chemical] and second order in [Chemical]. The activation energy of the reaction is [Value] kcal/mole.

The fourth part of the work was devoted to the study of the reaction of [Chemical] with [Chemical]. It was found that the reaction proceeds through a series of steps, the first of which is the formation of a complex. The rate of reaction is first order in [Chemical] and second order in [Chemical]. The activation energy of the reaction is [Value] kcal/mole.

The fifth part of the work was devoted to the study of the reaction of [Chemical] with [Chemical]. It was found that the reaction proceeds through a series of steps, the first of which is the formation of a complex. The rate of reaction is first order in [Chemical] and second order in [Chemical]. The activation energy of the reaction is [Value] kcal/mole.

The work was done in the laboratory of Professor [Name] during the year 1954. The work was supported by the National Science Foundation and the University of Chicago.

The following is a list of the publications resulting from this work:

- [Author], [Title], [Journal], [Year].
- [Author], [Title], [Journal], [Year].
- [Author], [Title], [Journal], [Year].