

NE MESAJO OIL CONSERVATION COMMISSION
Santa Fe, New Mexico

RECEIVED
MAY 5 1937

MISCELLANEOUS REPORTS ON WELLS

Submit this report in triplicate to the Oil Conservation Commission or its proper agent within ten days after the work specified is completed. It should be signed and sworn to before a notary public for reports on beginning drilling operations, results of shooting well, results of test of casing shut-off, result of plugging of well, and other important operations, even though the work was witnessed by an agent of the Commission. Reports on minor operations need not be signed and sworn to before a notary public. See additional instructions in the Rules and Regulations of the Commission.

Indicate nature of report by checking below:

REPORT ON BEGINNING DRILLING OPERATIONS		REPORT ON REPAIRING WELL	
REPORT ON RESULT OF SHOOTING OR CHEMICAL TREATMENT OF WELL		REPORT ON PULLING OR OTHERWISE ALTERING CASING	
REPORT ON RESULT OF TEST OF CASING SHUT-OFF	X	REPORT ON DEEPENING WELL	
REPORT ON RESULT OF PLUGGING OF WELL			

Monument, New Mexico

May 1, 1937

Place

Date

OIL CONSERVATION COMMISSION,
SANTA FE, NEW MEXICO.

Gentlemen:

Following is a report on the work done and the results obtained under the heading noted above at the _____
Amerada Petroleum Corporation State "Z" Well No. 1 in the _____
Company or Operator Lease
NE 1/4 NE 1/4 of Sec. 13, T. 20, R. 36, N. M. P. M.,
Monument Field, Lea County.

The dates of this work were as follows: _____
Notice of intention to do the work was [~~was not~~] submitted on Form C-102 on April 29, 1937 19____
and approval of the proposed plan was [was not] obtained. (Cross out incorrect words.)

DETAILED ACCOUNT OF WORK DONE AND RESULTS OBTAINED

8-5/8" 32# 8-Thd. New Seamless casing was set in this well at 2360' and cemented by the Halliburton Method with 600 sacks.

Casing and fittings were tested with 1200# pump pressure and allowed to stand undisturbed for thirty minutes. No drop in pressure resulted so the cement was then drilled out of the casing and the same test of 1200# pump pressure was again applied and allowed to stand undisturbed for thirty minutes. No drop in pressure resulted so the drilling was then resumed.

DUPLICATE

Witnessed by L.E. Stewart Name Noble Drilling Co. Company Toolpusher Title

Subscribed and sworn before me this _____
1 day of May, 1937
Lewis A. Deane
Notary Public
My commission expires Dec. 21, 1940

I hereby swear or affirm that the information given above is true and correct.

Name J. A. Sturkey
Position Sup't.
Representing Amerada Petroleum Corporation Company or Operator
Address Monument, New Mexico

Remarks:

Guy Shepard
Inspector Name
Oil Conservation Commission
SANTA FE, NEW MEX. Title

MAY 5 1937

THE UNIVERSITY OF CHICAGO

DEPARTMENT OF CHEMISTRY

1. The first part of the experiment involves the preparation of a standard solution of sodium hydroxide. This is done by weighing a precise amount of sodium hydroxide and dissolving it in a known volume of distilled water in a volumetric flask.

2. The second part of the experiment is the titration of a weak acid with the standard sodium hydroxide solution.

The weak acid is placed in a flask with an indicator. The standard sodium hydroxide solution is added from a burette until the color of the indicator changes, indicating the endpoint of the titration.

The volume of sodium hydroxide solution used is recorded, and the concentration of the weak acid is calculated using the stoichiometry of the reaction.

The results of the titration are compared with the theoretical values to determine the accuracy of the experiment.

3. The third part of the experiment involves the determination of the molar mass of a solid compound. This is done by measuring the mass of a known volume of a solution of the compound and comparing it to the mass of a known volume of a standard solution.

The molar mass is calculated using the relationship between mass, volume, and molarity.

The results are compared with the literature value to determine the accuracy of the experiment.

4. The fourth part of the experiment involves the determination of the concentration of a solution of a weak acid.

This is done by measuring the pH of the solution and using the relationship between pH and the concentration of the weak acid.

The concentration is calculated using the Henderson-Hasselbalch equation and the known pK_a of the weak acid.

5. The fifth part of the experiment involves the determination of the molar mass of a solid compound using gravimetric analysis.

This is done by weighing a known amount of the compound and heating it to a constant weight. The mass of the residue is compared to the mass of the original compound to determine the molar mass.

The results are compared with the literature value to determine the accuracy of the experiment.

6. The sixth part of the experiment involves the determination of the concentration of a solution of a weak acid using a colorimetric method.

This is done by measuring the absorbance of the solution at a specific wavelength and using the relationship between absorbance and concentration.

The concentration is calculated using Beer's Law and the known molar absorptivity of the weak acid.

7. The seventh part of the experiment involves the determination of the molar mass of a solid compound using a gravimetric method.

This is done by weighing a known amount of the compound and heating it to a constant weight. The mass of the residue is compared to the mass of the original compound to determine the molar mass.

The results are compared with the literature value to determine the accuracy of the experiment.

8. The eighth part of the experiment involves the determination of the concentration of a solution of a weak acid using a titrimetric method.

This is done by measuring the volume of a standard solution of a strong base required to titrate a known volume of the weak acid solution.

The concentration is calculated using the stoichiometry of the reaction and the known concentration of the standard solution.

9. The ninth part of the experiment involves the determination of the molar mass of a solid compound using a gravimetric method.

This is done by weighing a known amount of the compound and heating it to a constant weight. The mass of the residue is compared to the mass of the original compound to determine the molar mass.

The results are compared with the literature value to determine the accuracy of the experiment.