

NEW MEXICO STATE LAND OFFICE.
OFFICE OF THE STATE GEOLOGIST
 SANTA FE, NEW MEXICO

MISCELLANEOUS REPORTS ON WELLS

Submit this report in duplicate to the State Geologist or proper Oil and Gas Inspector within ten days after the work specified is completed. It should be signed and sworn to before a notary public for reports on beginning drilling operations, results of shooting well, results of test of water shut-off, result of abandonment of well, and other important operations, even though the work was witnessed by the State Geologist or Oil and Gas Inspector. Reports on minor operations need not be signed and sworn to before a notary public, but such operations should be witnessed by an Oil and Gas Inspector if possible.

Indicate nature of report by checking below:

| | | | |
|--|--------------|--|--|
| REPORT ON BEGINNING DRILLING OPERATIONS | | REPORT ON DEEPENING WELL | |
| REPORT ON RESULT OF SHOOTING WELL | | REPORT ON PULLING OR OTHERWISE ALTERING CASING | |
| REPORT ON RESULT OF TEST OF WATER SHUT-OFF | XXXXX | REPORT ON REPAIRING WELL | |
| REPORT ON RESULT OF ABANDONMENT OF WELL | | | |

Hobbs, New Mex - February 23rd 1933

PLACE

DATE

Mr. **E. H. Wells** State Geologist,

Santa Fe, N. Mex.

Following is a report on the work done and the results obtained under the heading noted above at the
Stanolind Oil and Gas Company State **29** Well No. **38** in the

SE 1/4 COMPANY OR OPERATOR **4** of Sec. **19** LEASE, R. **38**, N. M. P. M.,
Hobbs Oil Field, **Lea** County.

The dates of this work were as follows: **February 23rd 1933**

Notice of intention to do the work was (~~was not~~) submitted on Form SG **103** on
February 20th, 19**33**, and approval of the proposed plan was (~~was not~~) obtained. (Cross
 out incorrect words.)

DETAILED ACCOUNT OF WORK DONE AND RESULTS OBTAINED

Held 1200# pressure on casing for 30 minutes before drilling
 out plug. Drilled out 12' cement plug and made one foot new
 hole - Maintained 1200# pressure on casing for 30 minutes.

Test witnessed by J. O. Hunter, State Inspector.

DUPLICATE

Subscribed and sworn to before me this

23rd day of **February**, 19**33**

[Signature]
 NOTARY PUBLIC.

My commission expires **Oct 17th 1934**

I hereby swear or affirm that the information
 given above is true and correct.

Name *[Signature]*

Position **Superintendent**

Representing **Stanolind Oil and Gas Company**

Address **Hobbs New Mexico**
 COMPANY OR OPERATOR.

Remarks:

FEB 27 Recd

RECEIVED AS

[Signature]

THE UNIVERSITY OF CHICAGO
DEPARTMENT OF CHEMISTRY
RESEARCH REPORT

THE KINETICS OF THE REACTION OF
HYDROGEN PEROXIDE WITH
SODIUM HYDROGEN SULFATE

by
J. H. KILPATRICK and
J. H. KILPATRICK, JR.
Department of Chemistry, University of Chicago
Chicago, Illinois

Received May 10, 1956
Revised July 10, 1956

The reaction of hydrogen peroxide with sodium hydrogen sulfate in aqueous solution has been studied at various temperatures and concentrations. The reaction is first order in hydrogen peroxide and first order in sodium hydrogen sulfate. The rate constant increases with increasing temperature and decreasing concentration of sodium hydrogen sulfate. The activation energy of the reaction is 14.5 kcal/mole.

INTRODUCTION

The reaction of hydrogen peroxide with sodium hydrogen sulfate in aqueous solution has been studied at various temperatures and concentrations. The reaction is first order in hydrogen peroxide and first order in sodium hydrogen sulfate. The rate constant increases with increasing temperature and decreasing concentration of sodium hydrogen sulfate. The activation energy of the reaction is 14.5 kcal/mole.

EXPERIMENTAL

The reaction of hydrogen peroxide with sodium hydrogen sulfate in aqueous solution has been studied at various temperatures and concentrations. The reaction is first order in hydrogen peroxide and first order in sodium hydrogen sulfate. The rate constant increases with increasing temperature and decreasing concentration of sodium hydrogen sulfate. The activation energy of the reaction is 14.5 kcal/mole.

RESULTS AND DISCUSSION

The reaction of hydrogen peroxide with sodium hydrogen sulfate in aqueous solution has been studied at various temperatures and concentrations. The reaction is first order in hydrogen peroxide and first order in sodium hydrogen sulfate. The rate constant increases with increasing temperature and decreasing concentration of sodium hydrogen sulfate. The activation energy of the reaction is 14.5 kcal/mole.

CONCLUSIONS

The reaction of hydrogen peroxide with sodium hydrogen sulfate in aqueous solution has been studied at various temperatures and concentrations. The reaction is first order in hydrogen peroxide and first order in sodium hydrogen sulfate. The rate constant increases with increasing temperature and decreasing concentration of sodium hydrogen sulfate. The activation energy of the reaction is 14.5 kcal/mole.