

(SUBMIT IN TRIPLICATE)

UNITED STATES
DEPARTMENT OF THE INTERIOR
GEOLOGICAL SURVEY

Land Office L. C.
661100
Lease No. H
Unit _____

SUNDRY NOTICES AND REPORTS ON WELLS

NOTICE OF INTENTION TO DRILL	<input checked="" type="checkbox"/>	SUBSEQUENT REPORT OF WATER SHUT-OFF	
NOTICE OF INTENTION TO CHANGE PLANS		SUBSEQUENT REPORT OF SHOOTING OR ACIDIZING	
NOTICE OF INTENTION TO TEST WATER SHUT-OFF		SUBSEQUENT REPORT OF ALTERING CASING	
NOTICE OF INTENTION TO RE-DRILL OR REPAIR WELL		SUBSEQUENT REPORT OF RE-DRILLING OR REPAIR	
NOTICE OF INTENTION TO SHOOT OR ACIDIZE		SUBSEQUENT REPORT OF ABANDONMENT	
NOTICE OF INTENTION TO PULL OR ALTER CASING		SUPPLEMENTARY WELL HISTORY	
NOTICE OF INTENTION TO ABANDON WELL			

(INDICATE ABOVE BY CHECK MARK NATURE OF REPORT, NOTICE, OR OTHER DATA)

Federal - Trigg Lease

January 17

1957

Well No. 1 is located 660 ft. from S line and 660 ft. from E line of sec. 3
SW/4 3 20-S 33-E _____
(1/4 Sec. and Sec. No.) (Twp.) (Range) (Meridian)
undesignated Lee County New Mexico
(Field) (County or Subdivision) (State or Territory)

The elevation of the derrick floor above sea level is 3558 ft. 6-01 3553

DETAILS OF WORK

(State names of and expected depths to objective sands; show sizes, weights, and lengths of proposed casings; indicate mudding jobs, cementing points, and all other important proposed work)

Formations:

Yates --- Approx. 1100'
 Seven Rivers - " 3250'
 Total Depth - " 3500'

Drillings

0 - TD Cable Tool
 Closser Drilling Co.

Casings: 13-3/8" " 60'
 10-3/4" " 600'
 8-5/8" " 1000'
 5-1/2" " TD

I understand that this plan of work must receive approval in writing by the Geological Survey before operations may be commenced.

Company McCrath & Smith

Address P. O. Box 8777

Midland, Texas

By Pomeroy Smith
 Pomeroy Smith
 Title Operator

THE UNIVERSITY OF CHICAGO
DEPARTMENT OF CHEMISTRY
5780 SOUTH CAMPUS DRIVE
CHICAGO, ILLINOIS 60637

ANALYTICAL CHEMISTRY

1. A sample of a solid compound was analyzed and found to contain 40.0% carbon, 6.7% hydrogen, and 53.3% oxygen. The molecular weight of the compound is 132.0 g/mol. Determine the empirical and molecular formulas of the compound.

2. A 0.1000 g sample of a pure metal was dissolved in 10.00 mL of 1.000 M HCl. The resulting solution was diluted to 100.0 mL. The absorbance of the solution at 420 nm was 0.450. Calculate the molar absorptivity of the metal ion.

3. A 0.2000 g sample of a pure metal was dissolved in 10.00 mL of 1.000 M HCl. The resulting solution was diluted to 100.0 mL. The absorbance of the solution at 420 nm was 0.450. Calculate the molar absorptivity of the metal ion.

4. A 0.2000 g sample of a pure metal was dissolved in 10.00 mL of 1.000 M HCl. The resulting solution was diluted to 100.0 mL. The absorbance of the solution at 420 nm was 0.450. Calculate the molar absorptivity of the metal ion.

5. A 0.2000 g sample of a pure metal was dissolved in 10.00 mL of 1.000 M HCl. The resulting solution was diluted to 100.0 mL. The absorbance of the solution at 420 nm was 0.450. Calculate the molar absorptivity of the metal ion.

6. A 0.2000 g sample of a pure metal was dissolved in 10.00 mL of 1.000 M HCl. The resulting solution was diluted to 100.0 mL. The absorbance of the solution at 420 nm was 0.450. Calculate the molar absorptivity of the metal ion.

7. A 0.2000 g sample of a pure metal was dissolved in 10.00 mL of 1.000 M HCl. The resulting solution was diluted to 100.0 mL. The absorbance of the solution at 420 nm was 0.450. Calculate the molar absorptivity of the metal ion.

8. A 0.2000 g sample of a pure metal was dissolved in 10.00 mL of 1.000 M HCl. The resulting solution was diluted to 100.0 mL. The absorbance of the solution at 420 nm was 0.450. Calculate the molar absorptivity of the metal ion.

9. A 0.2000 g sample of a pure metal was dissolved in 10.00 mL of 1.000 M HCl. The resulting solution was diluted to 100.0 mL. The absorbance of the solution at 420 nm was 0.450. Calculate the molar absorptivity of the metal ion.

10. A 0.2000 g sample of a pure metal was dissolved in 10.00 mL of 1.000 M HCl. The resulting solution was diluted to 100.0 mL. The absorbance of the solution at 420 nm was 0.450. Calculate the molar absorptivity of the metal ion.

11. A 0.2000 g sample of a pure metal was dissolved in 10.00 mL of 1.000 M HCl. The resulting solution was diluted to 100.0 mL. The absorbance of the solution at 420 nm was 0.450. Calculate the molar absorptivity of the metal ion.

12. A 0.2000 g sample of a pure metal was dissolved in 10.00 mL of 1.000 M HCl. The resulting solution was diluted to 100.0 mL. The absorbance of the solution at 420 nm was 0.450. Calculate the molar absorptivity of the metal ion.

NEW MEXICO
OIL CONSERVATION COMMISSION

Form C-128

Well Location and/or Gas Proration Plat

Date 1-16-57

Operator McGRATH AND SMITH Lease FEO. TRIGG

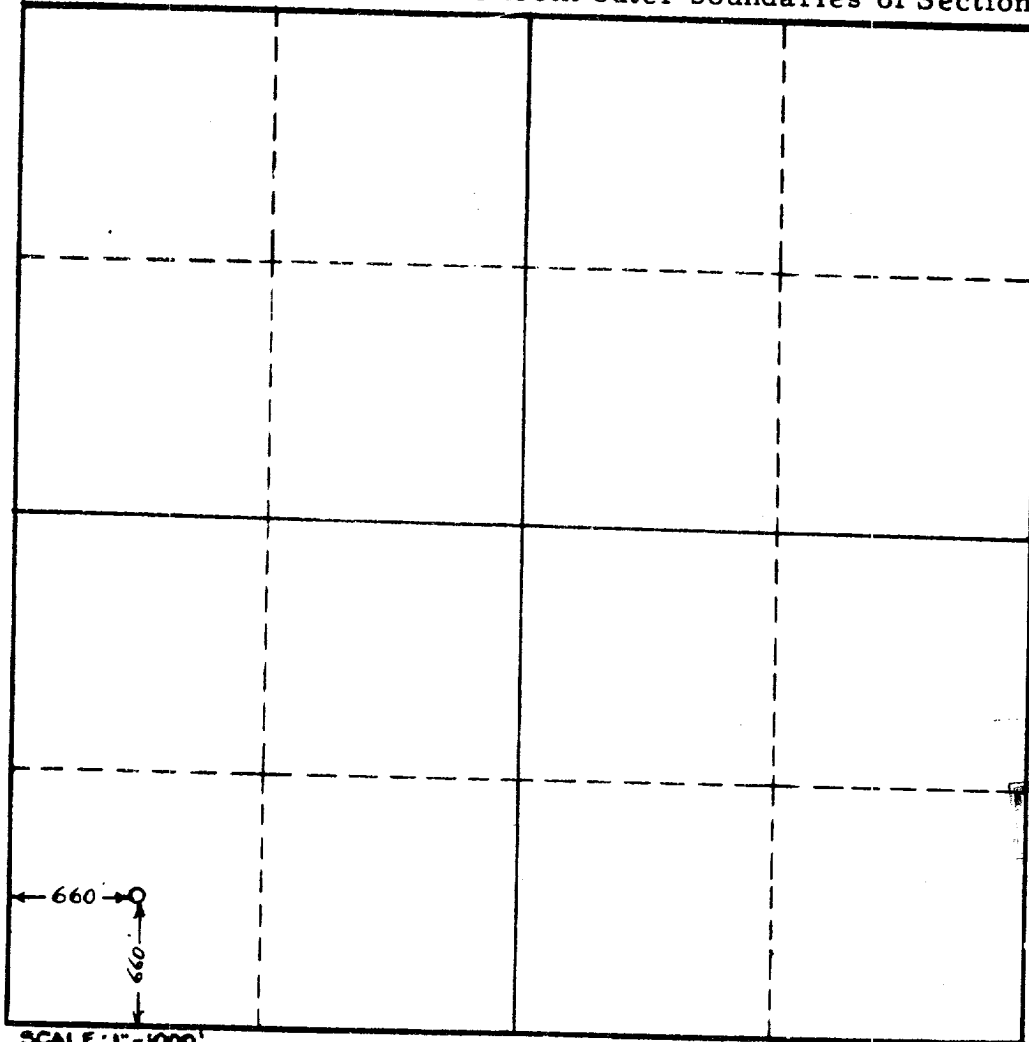
Well No. 1 Section 3 Township 20-S Range 33-E NMPM

Located 660 Feet From SOUTH Line, 660 Feet From WEST Line,

LEA County, New Mexico. G. L. Elevation _____

Name of Producing Formation _____ Pool _____ Dedicated Acreage _____

(Note: All distances must be from outer boundaries of Section)



SCALE: 1"=1000'

1. Is this Well a Dual Comp. ? Yes ___ No ___

2. If the answer to Question 1 is yes, are there any other dually completed wells within the dedicated acreage? Yes ___ No ___

This is to certify that the above plat was prepared from field notes of actual surveys made by me or under my supervision and that the same are true and correct to the best of my knowledge and belief.

Name _____
Position _____
Representing _____
Address _____

Date Surveyed 1-16-57
John W. West
Registered Professional Engineer and/or
Land Surveyor