

NEW MEXICO OIL CONSERVATION COMMISSION
Santa Fe, New Mexico

MISCELLANEOUS NOTICES

RECEIVED
JUN 16 1937

Submit this notice in triplicate to the Oil Conservation Commission or its proper agent before the work specified is to begin. A copy will be returned to the sender on which will be given the approval, with any modifications considered advisable, or the rejection by the Commission or its agent, of the plan submitted. The plan as approved should be followed, and work should not begin until approval is obtained. See additional instructions in the Rules and Regulations of the Commission.

Indicate nature of notice by checking below:

NOTICE OF INTENTION TO TEST CASING SHUT-OFF	10 3/4"	NOTICE OF INTENTION TO SHOOT OR CHEMICALLY TREAT WELL	
NOTICE OF INTENTION TO CHANGE PLANS		NOTICE OF INTENTION TO PULL OR OTHERWISE ALTER CASING	
NOTICE OF INTENTION TO REPAIR WELL		NOTICE OF INTENTION TO PLUG WELL	
NOTICE OF INTENTION TO DEEPEN WELL			

Hobbs, New Mexico June 14th 1937.

Place

Date

OIL CONSERVATION COMMISSION,

Santa Fe, New Mexico.

Gentlemen:

Following is a notice of intention to do certain work as described below at the

Gulf Oil Corp - Gypsy Division - R.R. Bell "F" Well No. #4 in SW/4
Company or Operator Lease
of Sec. 36, T. 20, R. 36, N. M. P. M., Eunice Field,
Lea. County.

FULL DETAILS OF PROPOSED PLAN OF WORK

FOLLOW INSTRUCTIONS IN THE RULES AND REGULATIONS OF THE COMMISSION

June 12th 1937 the 10 3/4" 32.75' 8thd new South Chester lapweld steel casing was cemented in Red Bed at 144' by the Halliburton Cementing process with 150 sacks cement

Propose to drill plug and test at 7 AM June 14th 1937.

DUPLICATE

Approved JUN 16 1937, 19____
except as follows:

OIL CONSERVATION COMMISSION,

By Guy Shepard R.M.
Title Gas Inspector

1.C.R.

GULF OIL CORPORATION
GYPSY DIVISION

By C.C. Cummings 298
Position District Supt.
Send communications regarding well to
Name C. C. Cummings.
Address Hobbs, New Mexico.

THE UNIVERSITY OF CHICAGO

CHICAGO, ILLINOIS

DEPARTMENT OF CHEMISTRY

For the purpose of this examination, the student is to be prepared to answer questions on the following topics: The structure and properties of the elements, the principles of chemical equilibrium, and the laws of thermodynamics. The student is to be able to apply these principles to the solution of problems in chemistry.

QUESTIONS

1. The following table gives the atomic weights of the elements. The student is to be able to calculate the molecular weight of a compound from these data. The student is to be able to calculate the percentage of an element in a compound from its molecular weight and atomic weight.

Element	Atomic Weight
Hydrogen	1.008
Helium	4.0026
Lithium	6.941
Boron	10.81
Carbon	12.011
Nitrogen	14.007
Oxygen	15.999
Fluorine	18.998
Neon	20.179
Sodium	22.990
Magnesium	24.305
Aluminum	26.981
Silicon	28.086
Phosphorus	30.974
Sulfur	32.06
Chlorine	35.453
Argon	39.948
Potassium	39.098
Calcium	40.078
Scandium	44.956
Titanium	47.88
Vanadium	50.942
Chromium	51.996
Manganese	54.938
Iron	55.845
Cobalt	58.933
Nickel	58.69
Copper	63.546
Zinc	65.38
Gallium	69.723
Germanium	72.63
Arsenic	74.922
Selenium	78.96
Bromine	79.904
Krypton	83.80
Rubidium	85.468
Strontium	87.62
Yttrium	88.906
Zirconium	91.224
Niobium	92.906
Molybdenum	95.94
Technetium	98.906
Ruthenium	101.07
Rhodium	102.91
Palladium	106.36
Silver	107.868
Cadmium	112.411
Indium	114.818
Mercury	200.59
Thallium	204.38
Lead	207.19
Bismuth	208.98
Polonium	209
Astatine	210
Radium	226
Actinium	227
Thorium	232
Protactinium	231
Uranium	238
Neptunium	237
Plutonium	244

2. The following table gives the boiling points of the elements. The student is to be able to calculate the heat of vaporization of a liquid from its boiling point and the heat of fusion of a solid from its melting point.

QUESTIONS

1. The following table gives the boiling points of the elements.

2. The following table gives the melting points of the elements.

3. The following table gives the heat of fusion of the elements. The student is to be able to calculate the heat of vaporization of a liquid from its boiling point and the heat of fusion of a solid from its melting point.

4. The following table gives the heat of vaporization of the elements. The student is to be able to calculate the heat of fusion of a solid from its melting point and the heat of vaporization of a liquid from its boiling point.

5. The following table gives the heat of fusion of the elements. The student is to be able to calculate the heat of vaporization of a liquid from its boiling point and the heat of fusion of a solid from its melting point.

QUESTIONS

THE UNIVERSITY OF CHICAGO

CHICAGO, ILLINOIS

For the purpose of this examination, the student is to be prepared to answer questions on the following topics: The structure and properties of the elements, the principles of chemical equilibrium, and the laws of thermodynamics. The student is to be able to apply these principles to the solution of problems in chemistry.

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2.

3.

4. The following table gives the boiling points of the elements.

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7. The following table gives the heat of vaporization of the elements. The student is to be able to calculate the heat of fusion of a solid from its melting point and the heat of vaporization of a liquid from its boiling point.

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