

May 21, 1974

MONUMENT ENERGY CORP. NO. 1 IVA SEDBERRY  
COMPLETION REPORT

Field: Wildcat

Location: 990' FN & 1650' FE Lines Section 25. T 17 N R 16 E  
San Miguel County, New Mexico

Elevation: 6572' G.L., 6584' K.B. T.D.: 8400' driller, 8411' log

Spudded: 4/20/74. Reached T.D.: 5/10/74 P & A: 5/15/74

Casing: 12-3/4" @ 86' G.L. w/150 sx.  
Drilled 9 1/2" hole to 2547'; 7-7/8" 2547' to T.D.

Cores: None

DST: No. 1 - 4620-4800:: Misrun

No. 2 - 4620-4800

15 min. preflow	142 psi
75 min. ISIP	157-171 psi
60 min. Flow	142 psi
60 min. FSIP	142-200 psi
H.P., initial & final	2414 psi
Weak blow air through 15 min. of flow period	
Rec. 30' drilling fluid, N/S	

Formation Tops:

Dakota	+6784 ( 300)	Glorieta	+4300 (2284)
Morrison	+6044 ( 540)	Upper Penn.	+3917 (2667)
Santa Rosa	+4729 (1855)	Lower Penn.	+2780 (3804)
Permian	+4544 (2040)	Top Sandia SS	+1954 (4630)

Geological Remarks:

Markers in subject well are virtually flat with those in the Hancock well, a short distance to the west. The top of Sandia is picked 20' higher than in the above well, but this is because of a stray bench of SS encountered @ 4630'. This bench was not present in the Hancock well.

## THEORY

The purpose of this experiment is to determine the rate of reaction between hydrogen peroxide and potassium iodide.

The reaction between hydrogen peroxide and potassium iodide is a redox reaction. The hydrogen peroxide is oxidized to water and the potassium iodide is reduced to potassium iodide.

The reaction is as follows:



The rate of reaction can be determined by measuring the volume of gas evolved over a period of time.

The rate of reaction is given by:

$$\text{Rate of reaction} = \frac{\text{Volume of gas evolved}}{\text{Time taken}}$$

The rate of reaction can be determined by measuring the volume of gas evolved over a period of time.

The rate of reaction is given by:

$$\text{Rate of reaction} = \frac{\text{Volume of gas evolved}}{\text{Time taken}}$$

The rate of reaction is given by:

$$\text{Rate of reaction} = \frac{\text{Volume of gas evolved}}{\text{Time taken}}$$

The rate of reaction is given by:

$$\text{Rate of reaction} = \frac{\text{Volume of gas evolved}}{\text{Time taken}}$$

The rate of reaction is given by:

The rate of reaction is given by:

The rate of reaction is given by:

The rate of reaction is given by:

The rate of reaction is given by: