

OIL CONSERVATION DIVISION

Drummer DD Artesia, NM

DISTRICT OFFICE #2

July thru August, 1980

NO. 2116 R

SUPPLEMENT TO THE OIL PRORATION SCHEDULE

DATE July 1, 1980

PURPOSE ALLOWABLE REVISION

Effective July 1, 1980, the allowables of the following Yates Pat. Corp. wells in the pools listed below are hereby revised as indicated.

ATOKA - YESO

Eddie NK #1-C-33-18-26 Increased to 19 BOPD

July Total - 589 bbls.
Aug. Total - 589 bbls.

Dowell MV #2-J-33-18-26 Increased to 16 BOPD

July Total - 496 bbls.
Aug. Total - 496 bbls.

FENASCO DRAW SA-YESO ASSOC

Mobil CI Fed. #8-I-6-19-25 Increased to 15 BOPD

July Total - 465 bbls.
Aug. Total - 465 bbls.

Irish Hills Yaso JE St. #3-E-12-19-24 Increased to 17 BOPD

July Total - 527 bbls.
Aug. Total - 527 bbls.

OIL CONSERVATION DIVISION

WAG:ar

Yates Pat. Corp.

HCO

W. H. [Signature]
DISTRICT SUPERVISOR

THEORY OF THE EXPERIMENT

The purpose of this experiment is to determine the value of the universal gas constant R by measuring the pressure, volume, and temperature of a gas.

The experiment is based on the ideal gas law:

$$PV = nRT$$

where P is the pressure, V is the volume, n is the number of moles, R is the universal gas constant, and T is the temperature.

In this experiment, the pressure P is measured using a manometer, the volume V is measured using a graduated cylinder, and the temperature T is measured using a thermometer.

The number of moles n is determined by the mass of the gas and its molar mass.

The universal gas constant R is then calculated using the ideal gas law.

The experimental setup consists of a manometer, a graduated cylinder, a thermometer, and a gas sample. The manometer is used to measure the pressure of the gas, the graduated cylinder is used to measure the volume of the gas, and the thermometer is used to measure the temperature of the gas.

The data collected from the experiment is used to calculate the value of the universal gas constant R .

The results of the experiment show that the value of the universal gas constant R is approximately $8.314 \text{ J/mol}\cdot\text{K}$.

The experiment was conducted under the following conditions:

- Pressure: $1.013 \times 10^5 \text{ Pa}$
- Volume: 0.0224 m^3
- Temperature: 273.15 K
- Number of moles: 1.0 mol

RESULTS AND DISCUSSION

Table 1

Table 2

Table 3

Table 4