

NEW MEXICO OIL CONSERVATION COMMISSION
Santa Fe, New Mexico

(Form C-104)
(Revised 7/1/52)

REQUEST FOR (OIL) - (GAS) ALLOWABLE

New Well
~~Recompletion~~

This form shall be submitted by the operator before an initial allowable will be assigned to any completed Oil or Gas well. Form C-104 is to be submitted in QUADRUPLICATE to the same District Office to which Form C-101 was sent. The allowable will be assigned effective 7:00 A.M. on date of completion or recompletion, provided this form is filed during calendar month of completion or recompletion. The completion date shall be that date in the case of an oil well when oil is delivered into the stock tanks. Gas must be reported on 15.025 psia at 60° Fahrenheit.

Farmington, N.M.

November 12, 1956

(Place)

(Date)

WE ARE HEREBY REQUESTING AN ALLOWABLE FOR A WELL KNOWN AS:

El Paso Natural Gas Co. Schwerdtfeger, Well No. 7-A, in SE 1/4 SE 1/4,
(Company or Operator) (Lease)

P, Sec. 31, T. 28N, R. 8W, NMPM, So. Blanco P.C. Ext. Pool
Unit Letter

San Juan County Date Spudded 8-25-56 Date Completed 10-19-56

Please indicate location:

D	C	B	A
E	F	G	H
L	K	J	I
M	N	O	P X

680'S, 852'E

Casing and Cementing Record

Size Feet Sax

9 5/8"	121'	100
5 1/2"	3000'	150
2"	2966'	---

Elevation 6712' Total Depth 3010' ~~Per C.O.~~ 2996'

Top oil/gas pay 2958' (Perf.) Name of Prod. Form Pictured Cliff

Casing Perforations: 2958-2980 or

Depth to Casing shoe of Prod. String 3010'

Natural Prod. Test BOPD

based on bbls. Oil in Hrs. Mins.

Test after acid or shot BOPD

Based on bbls. Oil in Hrs. Mins.

Gas Well Potential A.C.F. - 1,522 MCF/D, Ch. Volume - 1,415 MCF/D.

Size choke in inches 3/4"

Date first oil run to tanks or gas to Transmission system: W/O Pipeline

Transporter taking Oil or Gas: El Paso Natural

Remarks:

I hereby certify that the information given above is true and complete to the best of my knowledge.

Approved NOV 14 1956, 19

OIL CONSERVATION COMMISSION

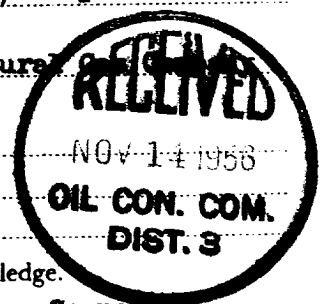
By: *A. F. Desjardis*
Title PETROLEUM ENGINEER DIST. NO. 3

El Paso Natural Gas Company
(Company or Operator)
By: *W. E. Schwerdtfeger*
(Signature)

Title: Petroleum Engineer
Send Communications regarding well to:

Name: E. J. Coel

Address: Box 997, Farmington, N.M.



OIL CONSERVATION COMMISSION

AZ 01-00079-00000-E

No. 6-1

2
1
1

Exposure time (h)	Cytosol fraction (%)
0	65
1	68
2	72
3	75
4	78
5	80
6	82
7	84
8	85
9	84
10	83
11	82
12	85
13	84
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87	10
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89	8
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91	6
92	5
93	4
94	3
95	2
96	1
97	0
98	0
99	0
100	0

Title	/
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DATE	1/1/19	TIME	10:00	LOCATION	1000
<p>1. The first part of the report is a description of the project. This includes the title, the objectives, the scope, and the methodology. The title is "The Effect of Temperature on the Rate of Reaction of Hydrogen Peroxide with Potassium Iodide". The objectives are to determine the effect of temperature on the rate of reaction and to determine the activation energy of the reaction. The scope is to study the reaction at temperatures between 10°C and 30°C. The methodology is to use a colorimetric method to measure the rate of reaction.</p>					
<p>2. The second part of the report is a description of the results. This includes the data collected, the graphs plotted, and the conclusions drawn. The data collected shows that the rate of reaction increases with temperature. The graphs plotted show a linear relationship between the natural logarithm of the rate constant and the inverse of the temperature. The conclusions drawn are that the reaction is first order with respect to hydrogen peroxide and second order with respect to potassium iodide. The activation energy of the reaction is 50 kJ/mol.</p>					
<p>3. The third part of the report is a discussion of the results. This includes a comparison of the results with the literature, a discussion of the sources of error, and a conclusion. The results are compared with the literature and found to be in good agreement. The sources of error are discussed and found to be small. The conclusion is that the reaction is first order with respect to hydrogen peroxide and second order with respect to potassium iodide. The activation energy of the reaction is 50 kJ/mol.</p>					
