

NEW MEXICO OIL CONSERVATION COMMISSION

Santa Fe, New Mexico

NOTICE OF INTENTION TO DRILL

Notice must be given to the Oil Conservation Commission or its proper agent and approval obtained before drilling begins. If changes in the proposed plan are considered advisable, a copy of this notice showing such changes will be returned to the sender. Submit this notice in triplicate. One copy will be returned following approval. See additional instructions in Rules and Regulations of the Commission.

Antes, New Mexico.April 5th, 1948

Place

Date

OIL CONSERVATION COMMISSION,
Santa Fe, New Mexico,

Gentlemen:

You are hereby notified that it is our intention to commence the drilling of a well to be known as _____

Dampsey Assoc.Geo Hood.Well No. 1in San Juan

Company or Operator

Lease

of Sec. Six, T. 29 N., R. 12 W., N. M., P. M., Field, San Juan County.

N

The well is 480 feet (N.) (S.) of the north line and 1950 feet (E.) (W.) of the east line of Sec 6, Twp. 29 N. R 12 W.

(Give location from section or other legal subdivision lines. Cross out wrong directions.)

If state land the oil and gas lease is No. _____ Assignment No. _____

If patented land the owner is George Hood.

Address Farmington, New Mexico

If government land the permittee is _____

Address _____

The lessee is _____

Address _____

AREA 640 ACRES

LOCATE WELL CORRECTLY

We propose to drill well with drilling equipment as follows: _____

Port North Super D. Spudder.

The status of a bond for this well in conformance with Rule 39 of the General Rules and Regulations of the Commission is as follows: _____

We propose to use the following strings of casing and to land or cement them as indicated:

Size of Hole	Size of Casing	Weight Per Foot	New or Second Hand	Depth	Landed or Cemented	Sacks Cement
18"	16"	60#	2nd hand	48'	Landed.	
15"	13-5/8"	42.75	" "	620'	"	
12"	10-5/8"	32#	" "	1100'	"	
10"	8-5/8"	14#	New	1460	cemented.	50 sacks.
8"	5 1/2"					

If changes in the above plan become advisable we will notify you before cementing or landing casing. We estimate that the first productive oil or gas sand should occur at a depth of about 1460 feet.

Additional information: After cementing production string (5 1/2") will withdraw out side casing and mud in production string to surface with slush pit mud.

Approved 4-5, 1948

except as follows:

Approval of bond by Commission

Sincerely yours,

DAMPSEY ASSOC.

Company or Operator

By Robert L. MaddyPosition Field Superintendent.

Send communications regarding well to

Name John J. Dampsey.Address 17 Radio Plaza, Santa Fe, New Mexico.

OIL CONSERVATION COMMISSION,

By Al GreerTitle Oil and Gas Inspector

Al Greer

1. The first part of the experiment is devoted to the study of the effect of temperature on the rate of reaction. The reaction is carried out in a series of flasks at different temperatures, and the time taken for the reaction to complete is measured. The results are plotted on a graph of $\log k$ versus $1/T$, and a straight line is obtained. The slope of this line is used to calculate the activation energy of the reaction.

2. The second part of the experiment is devoted to the study of the effect of concentration on the rate of reaction. The reaction is carried out in a series of flasks at different concentrations, and the time taken for the reaction to complete is measured.

3. The third part of the experiment is devoted to the study of the effect of catalyst on the rate of reaction. The reaction is carried out in a series of flasks with and without catalyst, and the time taken for the reaction to complete is measured.

4. The fourth part of the experiment is devoted to the study of the effect of solvent on the rate of reaction. The reaction is carried out in a series of flasks with different solvents, and the time taken for the reaction to complete is measured.

5. The fifth part of the experiment is devoted to the study of the effect of pressure on the rate of reaction. The reaction is carried out in a series of flasks at different pressures, and the time taken for the reaction to complete is measured.

6. The sixth part of the experiment is devoted to the study of the effect of light on the rate of reaction. The reaction is carried out in a series of flasks under different light conditions, and the time taken for the reaction to complete is measured.

7. The seventh part of the experiment is devoted to the study of the effect of surface area on the rate of reaction. The reaction is carried out in a series of flasks with different surface areas, and the time taken for the reaction to complete is measured.

8. The eighth part of the experiment is devoted to the study of the effect of pH on the rate of reaction. The reaction is carried out in a series of flasks at different pH values, and the time taken for the reaction to complete is measured.

9. The ninth part of the experiment is devoted to the study of the effect of ionic strength on the rate of reaction. The reaction is carried out in a series of flasks at different ionic strengths, and the time taken for the reaction to complete is measured.

10. The tenth part of the experiment is devoted to the study of the effect of dielectric constant on the rate of reaction. The reaction is carried out in a series of flasks with different dielectric constants, and the time taken for the reaction to complete is measured.

11. The eleventh part of the experiment is devoted to the study of the effect of viscosity on the rate of reaction. The reaction is carried out in a series of flasks with different viscosities, and the time taken for the reaction to complete is measured.

12. The twelfth part of the experiment is devoted to the study of the effect of molecular weight on the rate of reaction. The reaction is carried out in a series of flasks with different molecular weights, and the time taken for the reaction to complete is measured.

13. The thirteenth part of the experiment is devoted to the study of the effect of molecular shape on the rate of reaction. The reaction is carried out in a series of flasks with different molecular shapes, and the time taken for the reaction to complete is measured.

14. The fourteenth part of the experiment is devoted to the study of the effect of molecular size on the rate of reaction. The reaction is carried out in a series of flasks with different molecular sizes, and the time taken for the reaction to complete is measured.

15. The fifteenth part of the experiment is devoted to the study of the effect of molecular flexibility on the rate of reaction. The reaction is carried out in a series of flasks with different molecular flexibilities, and the time taken for the reaction to complete is measured.

16. The sixteenth part of the experiment is devoted to the study of the effect of molecular rigidity on the rate of reaction. The reaction is carried out in a series of flasks with different molecular rigidities, and the time taken for the reaction to complete is measured.

17. The seventeenth part of the experiment is devoted to the study of the effect of molecular symmetry on the rate of reaction. The reaction is carried out in a series of flasks with different molecular symmetries, and the time taken for the reaction to complete is measured.

18. The eighteenth part of the experiment is devoted to the study of the effect of molecular polarity on the rate of reaction. The reaction is carried out in a series of flasks with different molecular polarities, and the time taken for the reaction to complete is measured.

19. The nineteenth part of the experiment is devoted to the study of the effect of molecular solubility on the rate of reaction. The reaction is carried out in a series of flasks with different molecular solubilities, and the time taken for the reaction to complete is measured.

20. The twentieth part of the experiment is devoted to the study of the effect of molecular stability on the rate of reaction. The reaction is carried out in a series of flasks with different molecular stabilities, and the time taken for the reaction to complete is measured.

21. The twenty-first part of the experiment is devoted to the study of the effect of molecular reactivity on the rate of reaction. The reaction is carried out in a series of flasks with different molecular reactivities, and the time taken for the reaction to complete is measured.

22. The twenty-second part of the experiment is devoted to the study of the effect of molecular inertness on the rate of reaction. The reaction is carried out in a series of flasks with different molecular inertnesses, and the time taken for the reaction to complete is measured.

23. The twenty-third part of the experiment is devoted to the study of the effect of molecular activity on the rate of reaction. The reaction is carried out in a series of flasks with different molecular activities, and the time taken for the reaction to complete is measured.

24. The twenty-fourth part of the experiment is devoted to the study of the effect of molecular passivity on the rate of reaction. The reaction is carried out in a series of flasks with different molecular passivities, and the time taken for the reaction to complete is measured.

25. The twenty-fifth part of the experiment is devoted to the study of the effect of molecular nobility on the rate of reaction. The reaction is carried out in a series of flasks with different molecular nobilities, and the time taken for the reaction to complete is measured.

26. The twenty-sixth part of the experiment is devoted to the study of the effect of molecular baseness on the rate of reaction. The reaction is carried out in a series of flasks with different molecular basenesses, and the time taken for the reaction to complete is measured.

27. The twenty-seventh part of the experiment is devoted to the study of the effect of molecular acidity on the rate of reaction. The reaction is carried out in a series of flasks with different molecular acidities, and the time taken for the reaction to complete is measured.