

Form 9-231a  
(Feb. 1951)

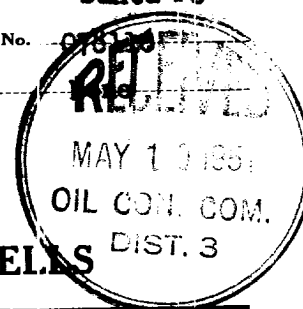
(SUBMIT IN TRIPLICATE)

UNITED STATES  
DEPARTMENT OF THE INTERIOR  
GEOLOGICAL SURVEY

Land Office: Santa Fe

Lease No. 078120

Unit: Relined



SUNDRY NOTICES AND REPORTS ON WELLS

NOTICE OF INTENTION TO DRILL.....	<input checked="" type="checkbox"/>	SUBSEQUENT REPORT OF WATER SHUT-OFF.....	
NOTICE OF INTENTION TO CHANGE PLANS.....		SUBSEQUENT REPORT OF SHOOTING OR ACIDIZING.....	
NOTICE OF INTENTION TO TEST WATER SHUT-OFF.....		SUBSEQUENT REPORT OF ALTERING CASING.....	
NOTICE OF INTENTION TO RE-DRILL OR REPAIR WELL.....		SUBSEQUENT REPORT OF REDRILLING OR REPAIR.....	
NOTICE OF INTENTION TO SHOOT OR ACIDIZE.....		SUBSEQUENT REPORT OF ABANDONMENT.....	
NOTICE OF INTENTION TO PULL OR ALTER CASING.....		SUPPLEMENTARY WELL HISTORY.....	
NOTICE OF INTENTION TO ABANDON WELL.....			

(INDICATE ABOVE BY CHECK MARK NATURE OF REPORT, NOTICE, OR OTHER DATA)

April 11, 1950

Florance-Federal, et al.  
Well No. 1 is located 990 ft. from NE line and 990 ft. from E line of sec. 21

SW 1/4 SEC. 21 30N 9W NMPM  
(1/4 Sec. and Sec. No.) (Twp.) (Range) (Meridian)

Blanco San Juan New Mexico  
(Field) (County or Subdivision) (State or Territory)

The elevation of the derrick floor above sea level is 6013 ft. gr.

DETAILS OF WORK

(State names of and expected depths to objective sands; show sizes, weights, and lengths of proposed casings; indicate mudding jobs, cementing points, and all other important proposed work)

This is to be a Mesa Verde test:

Hole will be drilled from surface to top of Cliff House sand with rotary tools. 7" casing will be set and cemented. Hole will be deepened through Point Lookout sand with cable tools. Pt. Lookout will be shot with approx. 70 lbs. nitroglycerene.

Approximate casing program will be:

100'	13-3/8"	40" seamless casing	cemented to surface
4200'	20 #	7" seamless casing	250 sacks cement

I understand that this plan of work must receive approval in writing by the Geological Survey before operations may be commenced.

Company Delhi Oil Corporation

Address 1315 Pacific Ave.

Dallas, Texas

By \_\_\_\_\_

Title Frank A. Schultze, Vice-President

1. The first part of the report is a general description of the project. It includes the title, the objectives, the scope, and the methodology. The title is "A Study of the Effect of Temperature on the Rate of Reaction of Hydrogen Peroxide with Potassium Iodide". The objectives are to determine the effect of temperature on the rate of reaction and to determine the activation energy of the reaction. The scope is to study the reaction at temperatures ranging from 10°C to 40°C. The methodology is to use the method of initial rates.

2. The second part of the report is a description of the experimental procedure. It includes the list of materials, the apparatus, and the steps of the experiment. The materials are hydrogen peroxide, potassium iodide, and a thermometer. The apparatus is a conical flask, a stop clock, and a water bath. The steps of the experiment are to prepare a series of solutions of different concentrations of hydrogen peroxide and potassium iodide, to measure the time taken for the reaction to occur at different temperatures, and to calculate the rate of reaction from the time taken.

3. The third part of the report is a description of the results of the experiment. It includes a table of the data, a graph of the rate of reaction against temperature, and a calculation of the activation energy. The table shows that the rate of reaction increases with temperature. The graph shows a linear relationship between the logarithm of the rate of reaction and the reciprocal of the absolute temperature. The activation energy is calculated to be 50 kJ/mol.

4. The fourth part of the report is a discussion of the results. It compares the results with the theoretical predictions and discusses the sources of error. The results are in good agreement with the theoretical predictions. The sources of error are the uncertainty in the measurement of time and the uncertainty in the measurement of temperature.

5. The fifth part of the report is a conclusion. It summarizes the findings of the experiment and states the overall conclusion. The findings are that the rate of reaction increases with temperature and that the activation energy of the reaction is 50 kJ/mol. The overall conclusion is that the method of initial rates is a reliable method for determining the rate of reaction.

6. The sixth part of the report is a list of references. It includes the names of the authors and the titles of the books and articles that were consulted during the preparation of the report. The references are: "Chemical Kinetics" by P. D. Bartlett, "The Rate of Reaction of Hydrogen Peroxide with Potassium Iodide" by J. H. Golding, and "The Effect of Temperature on the Rate of Reaction of Hydrogen Peroxide with Potassium Iodide" by A. R. Fieser.