

Oil Reports and Gas Services, Inc.

P. O. BOX 763 HOBBS, NEW MEXICO 88240

August 2, 1977

New Mexico Oil Conservation Commission
P. O. Box 2088
Santa Fe, New Mexico 87501

Gentlemen:

Application is hereby made on behalf of Dallas McCasland for administrative approval of an 80-acre non-standard Jalmat Gas Proration Unit to comprise the W/2 SW/4 of Section 16, Township 23 South, Range 36 East, to be dedicated to the Aztec State No. 1, located 990 feet from the South line and 660 feet from the West line of said Section 16.

The subject well has produced since completion as a Jalmat oil well. By letter dated July 27, 1977, the Oil Conservation Commission has advised that it will be reclassified as a gas well due to excessive gas-oil ratio.

Attached is a plat showing the Aztec State lease and the offset operators. All offset operators have been furnished a copy of this application and plat by certified mail.

Yours very truly,

OIL REPORTS & GAS SERVICES, INC.

OFFIC. SIGNED BY DONNA HOLLER

(Mrs.) Donna Holler

DH/kca

cc: Atlantic Richfield Company
P. O. Box 1610
Midland, Texas 79701

Highland Production Company
P. O. Box 6725
Odessa, Texas 79762

Maralo, Inc.
P. O. Box 832
Midland, Texas 79701

Dallas McCasland
P. O. Box 206
Eunice, New Mexico 88231

Oil Conservation Commission
P. O. Box 1960
Hobbs, New Mexico 88240

Texasco, Inc.
P. O. Box 3109
Midland, Texas 79701

Texas Pacific Oil Company, Inc.
P. O. Box 4067
Midland, Texas 79701

THE UNIVERSITY OF CHICAGO
DEPARTMENT OF CHEMISTRY

1. The first part of the experiment is to determine the concentration of the solution. This is done by measuring the volume of the solution and the mass of the solute. The concentration is then calculated as the mass of the solute divided by the volume of the solution.

2. The second part of the experiment is to determine the molar mass of the solute. This is done by measuring the change in boiling point of the solvent when a known mass of solute is added. The molar mass is then calculated as the mass of the solute divided by the change in boiling point.

3. The third part of the experiment is to determine the van't Hoff factor of the solute. This is done by measuring the change in boiling point of the solvent when a known mass of solute is added. The van't Hoff factor is then calculated as the change in boiling point divided by the change in boiling point for a non-electrolyte.

4. The fourth part of the experiment is to determine the dissociation constant of the solute. This is done by measuring the change in boiling point of the solvent when a known mass of solute is added. The dissociation constant is then calculated as the change in boiling point divided by the change in boiling point for a non-electrolyte.

5. The fifth part of the experiment is to determine the degree of dissociation of the solute. This is done by measuring the change in boiling point of the solvent when a known mass of solute is added. The degree of dissociation is then calculated as the change in boiling point divided by the change in boiling point for a non-electrolyte.

6. The sixth part of the experiment is to determine the equilibrium constant of the reaction. This is done by measuring the change in boiling point of the solvent when a known mass of solute is added. The equilibrium constant is then calculated as the change in boiling point divided by the change in boiling point for a non-electrolyte.

7. The seventh part of the experiment is to determine the Gibbs free energy of the reaction. This is done by measuring the change in boiling point of the solvent when a known mass of solute is added. The Gibbs free energy is then calculated as the change in boiling point divided by the change in boiling point for a non-electrolyte.

8. The eighth part of the experiment is to determine the entropy of the reaction. This is done by measuring the change in boiling point of the solvent when a known mass of solute is added. The entropy is then calculated as the change in boiling point divided by the change in boiling point for a non-electrolyte.

9. The ninth part of the experiment is to determine the enthalpy of the reaction. This is done by measuring the change in boiling point of the solvent when a known mass of solute is added. The enthalpy is then calculated as the change in boiling point divided by the change in boiling point for a non-electrolyte.

10. The tenth part of the experiment is to determine the free energy of the reaction. This is done by measuring the change in boiling point of the solvent when a known mass of solute is added. The free energy is then calculated as the change in boiling point divided by the change in boiling point for a non-electrolyte.

