

NEW MEXICO OIL CONSERVATION COMM.

Santa Fe, New Mexico

MISCELLANEOUS REPORTS ON WELLS

Submit this report in triplicate to the Oil Conservation Commission or its proper agent within ten days after the work specified is completed. It should be signed and sworn to before a notary public for reports on beginning drilling operations, results of shooting well, results of test of casing shut-off, result of plugging of well, and other important operations, even though the work was witnessed by an agent of the Commission. Reports on minor operations need not be signed and sworn to before a notary public. See additional instructions in the Rules and Regulations of the Commission.

Indicate nature of report by checking below:

REPORT ON BEGINNING DRILLING OPERATIONS		REPORT ON REPAIRING WELL	
REPORT ON RESULT OF SHOOTING OR CHEMICAL TREATMENT OF WELL		REPORT ON PULLING OR OTHERWISE ALTERING CASING	
REPORT ON RESULT OF TEST OF CASING SHUT-OFF		REPORT ON DEEPENING WELL	
REPORT ON RESULT OF PLUGGING OF WELL	<input checked="" type="checkbox"/>		

Artesia, New Mexico

August 10, 1943

Place

Date

**OIL CONSERVATION COMMISSION,
SANTA FE, NEW MEXICO.**

Gentlemen:

Following is a report on the work done and the results obtained under the heading noted above at the _____

Cunningham & Ployhar

State

Well No. 1

in the

Company or Operator

Lease

11 22 22

of Sec. 30

T. 17

R. 28

N. M. P. M.,

Red Lake

Field,

Eddy

County.

The dates of this work were as follows: July 14, 1943

Notice of intention to do the work was ~~(was not)~~ submitted on Form C-102 on June 11, 1943

and approval of the proposed plan was ~~(was not)~~ obtained. (Cross out incorrect words.)

DETAILED ACCOUNT OF WORK DONE AND RESULTS OBTAINED

We plugged the above well as follows:

Filled hole from 420 (TD) to 380 with cement; knocked off and pulled 280' of 7" cs; then filled with mud to 280 and set bridge and cement plug; then filled with mud to surface and cemented regulation marker.

Witnessed by Walter Solt

Self

Name

Company

Title

Subscribed and sworn before me this _____

I hereby swear or affirm that the information given above is true and correct.

_____ day of _____, 19_____

Name W. D. Cunningham (Signed)

Position Partner

Notary Public

Representing Cunningham & Ployhar
Company or Operator

My commission expires _____

Address Box 163, Artesia, New Mexico

Remarks:

Roy Yarbrough (Signed)

Name

Oil & Gas Inspector

Title

THE UNIVERSITY OF CHICAGO
DEPARTMENT OF CHEMISTRY
PHYSICAL CHEMISTRY

1. The first part of the experiment is devoted to the study of the temperature dependence of the rate constant for the reaction of hydrogen peroxide with iodide ions in the presence of ceric ions as a catalyst. The reaction is carried out in a series of solutions of known concentrations of the reactants and the catalyst, and the rate of reaction is measured by the appearance of a color due to the formation of iodine. The rate constant is determined from the slope of a plot of the logarithm of the rate constant versus the inverse of the absolute temperature.

2. The second part of the experiment is devoted to the study of the temperature dependence of the rate constant for the reaction of hydrogen peroxide with iodide ions in the presence of ceric ions as a catalyst. The reaction is carried out in a series of solutions of known concentrations of the reactants and the catalyst, and the rate of reaction is measured by the appearance of a color due to the formation of iodine. The rate constant is determined from the slope of a plot of the logarithm of the rate constant versus the inverse of the absolute temperature.

3. The third part of the experiment is devoted to the study of the temperature dependence of the rate constant for the reaction of hydrogen peroxide with iodide ions in the presence of ceric ions as a catalyst. The reaction is carried out in a series of solutions of known concentrations of the reactants and the catalyst, and the rate of reaction is measured by the appearance of a color due to the formation of iodine. The rate constant is determined from the slope of a plot of the logarithm of the rate constant versus the inverse of the absolute temperature.

4. The fourth part of the experiment is devoted to the study of the temperature dependence of the rate constant for the reaction of hydrogen peroxide with iodide ions in the presence of ceric ions as a catalyst. The reaction is carried out in a series of solutions of known concentrations of the reactants and the catalyst, and the rate of reaction is measured by the appearance of a color due to the formation of iodine. The rate constant is determined from the slope of a plot of the logarithm of the rate constant versus the inverse of the absolute temperature.

5. The fifth part of the experiment is devoted to the study of the temperature dependence of the rate constant for the reaction of hydrogen peroxide with iodide ions in the presence of ceric ions as a catalyst. The reaction is carried out in a series of solutions of known concentrations of the reactants and the catalyst, and the rate of reaction is measured by the appearance of a color due to the formation of iodine. The rate constant is determined from the slope of a plot of the logarithm of the rate constant versus the inverse of the absolute temperature.