



(SUBMIT IN TRIPLICATE)

UNITED STATES
DEPARTMENT OF THE INTERIOR
GEOLOGICAL SURVEY

Office Las Cruces

License No. L. G. 097152 B

Unit Las Cruces

SUNDRY NOTICES AND REPORTS ON WELLS

NOTICE OF INTENTION TO DRILL.....	<input checked="" type="checkbox"/>	SUBSEQUENT REPORT OF WATER SHUT-OFF.....	
NOTICE OF INTENTION TO CHANGE PLANS.....		SUBSEQUENT REPORT OF SHOOTING OR ACIDIZING.....	
NOTICE OF INTENTION TO TEST WATER SHUT-OFF.....		SUBSEQUENT REPORT OF ALTERING CASING.....	
NOTICE OF INTENTION TO RE-DRILL OR REPAIR WELL.....		SUBSEQUENT REPORT OF RE-DRILLING OR REPAIR.....	
NOTICE OF INTENTION TO SHOOT OR ACIDIZE.....		SUBSEQUENT REPORT OF ABANDONMENT.....	
NOTICE OF INTENTION TO PULL OR ALTER CASING.....		SUPPLEMENTARY WELL HISTORY.....	
NOTICE OF INTENTION TO ABANDON WELL.....			

(INDICATE ABOVE BY CHECK MARK NATURE OF REPORT, NOTICE, OR OTHER DATA)

John B **November 19,** , 19**57**

Well No. **4** is located **2300** ft. from **18N** line and **1650** ft. from **12W** line of sec. **24**

17S **12E** **10W**
(Twp.) (Range) (Meridian)

Maljamar **Don** **New Mexico**
(Field) (County or Subdivision) (State or Territory)

The elevation of the derrick floor above sea level is **4063** ft.

DETAILS OF WORK

(State names of and expected depths to objective sands; show sizes, weights, and lengths of proposed casings; indicate mudding jobs, cementing points, and all other important proposed work)

We propose to drill at the above location with rotary tools to approximately 4350 feet. Surface casing to consist of 8 5/8 inch casing cemented at 150 feet with 150 sacks regular cement, circulated. Oil string to consist of 5 1/2 inch casing cemented at total depth with 400 sacks cement. Perforations to be opposite San Andres as determined by radioactive log. Well to be tested natural and if necessary treated with 500 gallon mud acid and 10,000 gallons Sandfree.

I understand that this plan of work must receive approval in writing by the Geological Survey before operations may be commenced.

Company Drilling & Exploration Company, Inc.

Address Box 2975, N

Hobbs, New Mexico

By [Signature]

Title Division Prod. Superintendent

THE UNIVERSITY OF CHICAGO
DEPARTMENT OF CHEMISTRY
PHYSICAL CHEMISTRY
LABORATORY

1. The first part of the experiment is devoted to the study of the temperature dependence of the rate constant for the reaction of hydrogen peroxide with iodide ions in the presence of ceric ions as a catalyst. The reaction is carried out in a series of solutions of known concentrations of the reactants and the products. The rate of reaction is measured by the appearance of a color due to the formation of iodine. The rate constant is determined from the initial rate of reaction and the concentrations of the reactants. The temperature dependence of the rate constant is studied by carrying out the reaction at several different temperatures. The activation energy of the reaction is determined from the Arrhenius plot of the rate constant versus the inverse of the absolute temperature.

2. The second part of the experiment is devoted to the study of the effect of the concentration of the reactants on the rate of reaction. The reaction is carried out in a series of solutions of known concentrations of the reactants and the products. The rate of reaction is measured by the appearance of a color due to the formation of iodine. The rate constant is determined from the initial rate of reaction and the concentrations of the reactants. The effect of the concentration of the reactants on the rate of reaction is studied by carrying out the reaction in solutions of different concentrations of the reactants. The order of reaction with respect to each reactant is determined from the initial rate of reaction and the concentrations of the reactants.

3. The third part of the experiment is devoted to the study of the effect of the concentration of the products on the rate of reaction. The reaction is carried out in a series of solutions of known concentrations of the reactants and the products. The rate of reaction is measured by the appearance of a color due to the formation of iodine. The rate constant is determined from the initial rate of reaction and the concentrations of the reactants. The effect of the concentration of the products on the rate of reaction is studied by carrying out the reaction in solutions of different concentrations of the products. The order of reaction with respect to each product is determined from the initial rate of reaction and the concentrations of the products.

4. The fourth part of the experiment is devoted to the study of the effect of the concentration of the catalyst on the rate of reaction. The reaction is carried out in a series of solutions of known concentrations of the reactants and the products. The rate of reaction is measured by the appearance of a color due to the formation of iodine. The rate constant is determined from the initial rate of reaction and the concentrations of the reactants. The effect of the concentration of the catalyst on the rate of reaction is studied by carrying out the reaction in solutions of different concentrations of the catalyst. The order of reaction with respect to the catalyst is determined from the initial rate of reaction and the concentrations of the catalyst.

5. The fifth part of the experiment is devoted to the study of the effect of the concentration of the reactants on the rate of reaction in the presence of a catalyst. The reaction is carried out in a series of solutions of known concentrations of the reactants and the products. The rate of reaction is measured by the appearance of a color due to the formation of iodine. The rate constant is determined from the initial rate of reaction and the concentrations of the reactants. The effect of the concentration of the reactants on the rate of reaction in the presence of a catalyst is studied by carrying out the reaction in solutions of different concentrations of the reactants and the catalyst. The order of reaction with respect to each reactant and the catalyst is determined from the initial rate of reaction and the concentrations of the reactants and the catalyst.