

N. M. OIL CONSERVATION COMMISSION

Santa Fe, New Mexico

MISCELLANEOUS NOTICES

Submit this notice in triplicate to the Oil Commission or its proper agent before the work specified is to begin. A copy will be returned to the sender on which will be given the approval, with any modifications considered advisable, or the rejection by the Commissioner or agent, of the plan submitted. The plan as approved should be followed, and work should not begin until approval is obtained. See additional instructions in the Rules and Regulations of the Commission.

Indicate nature of notice by checking below:

NOTICE OF INTENTION TO TEST CASING SHUT-OFF	X	NOTICE OF INTENTION TO SHOOT OR CHEMICALLY TREAT WELL
NOTICE OF INTENTION TO CHANGE PLANS		NOTICE OF INTENTION TO PULL OR OTHERWISE ALTER CASING
NOTICE OF INTENTION TO REPAIR WELL		
NOTICE OF INTENTION TO DEEPEN WELL		NOTICE OF INTENTION TO PLUG WELL

Lovington, New Mexico. April 29th, 1939.

Place

Date

OIL CONSERVATION COMMISSION,
Santa Fe, New Mexico.

Gentlemen:

DUPLICATE

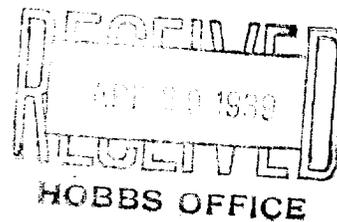
Following is a notice of intention to do certain work as described below at the

Magnolia Petroleum Co., State "N" Well No. 1 in NE 1/4 NE 1/4
 Company or Operator Lease
 of Sec. 34, T. 16S, R. 34E, N. M. P. M., Vacuum Field,
 Lea County.

FULL DETAILS OF PROPOSED PLAN OF WORK

FOLLOW INSTRUCTIONS IN THE RULES AND REGULATIONS OF THE COMMISSION

1665'
 Cemented 9-5/8" Casing @ 1665' with 200 Sx Cement & 7 Aquagel
 will Drill Plug and Test Casing Shut-off.



Approved Apr 29 1939, 19

MAGNOLIA PETROLEUM COMPANY

Company or Operator

By R. H. Alexander

Position Supt.

Send communications regarding well to

Name Magnolia Petroleum Co.,

Address Box 68, Lovington, New Mexico.

OIL CONSERVATION COMMISSION,

By Ray Yarrbraugh

Title OIL & GAS INSPECTOR

THE UNIVERSITY OF CHICAGO
DEPARTMENT OF CHEMISTRY
PHYSICAL CHEMISTRY

1. The first part of the experiment is devoted to the study of the temperature dependence of the rate constant for the reaction of hydrogen peroxide with iodide ions in the presence of ceric ions as a catalyst. The reaction is carried out in a series of solutions of known concentrations of the reactants and the catalyst. The rate of the reaction is measured by the appearance of a color due to the formation of iodine. The rate constant is determined from the initial rate of the reaction and the concentrations of the reactants. The temperature dependence of the rate constant is studied by carrying out the reaction at several different temperatures. The activation energy of the reaction is determined from the Arrhenius plot of the rate constant versus the inverse of the absolute temperature.

2. The second part of the experiment is devoted to the study of the effect of the concentration of the reactants on the rate of the reaction. The reaction is carried out in a series of solutions of known concentrations of the reactants and the catalyst. The rate of the reaction is measured by the appearance of a color due to the formation of iodine. The rate constant is determined from the initial rate of the reaction and the concentrations of the reactants. The effect of the concentration of the reactants on the rate of the reaction is studied by carrying out the reaction in solutions of different concentrations of the reactants. The order of the reaction with respect to each reactant is determined from the log-log plot of the rate of the reaction versus the concentration of the reactant.

3. The third part of the experiment is devoted to the study of the effect of the concentration of the catalyst on the rate of the reaction. The reaction is carried out in a series of solutions of known concentrations of the reactants and catalyst. The rate of the reaction is measured by the appearance of a color due to the formation of iodine. The rate constant is determined from the initial rate of the reaction and the concentrations of the reactants. The effect of the concentration of the catalyst on the rate of the reaction is studied by carrying out the reaction in solutions of different concentrations of the catalyst. The order of the reaction with respect to the catalyst is determined from the log-log plot of the rate of the reaction versus the concentration of the catalyst.