

NEW MEXICO OIL CONSERVATION COMMISSION

Santa Fe, New Mexico

REQUEST FOR PERMISSION TO CONNECT WITH PIPE LINE

THIS REQUEST SHOULD BE SUBMITTED IN TRIPLICATE. See instructions in the Rules and Regulations of the Commission.

Monument, New Mexico. August 21, 1939.

Place

Date

OIL CONSERVATION COMMISSION,
Santa Fe, New Mexico.

DUPLICATE

Gentlemen:

Permission is requested to connect Amerada Petroleum Corporation State "E"
 Company or Operator Lease
 Wells No. 1 and 2 in NW 1/4 of Sec. 24 T. 18s R. 37e N. M. P. M.,

Hobbs Field, Lea County, with the pipe line of the
Humble Pipe Line Company Houston, Texas.

Pipe Line Co.

State

Address

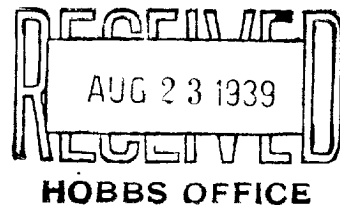
Status of land (State, Government or privately owned)

Location of tank battery 1320' FNL and 660' FNL Section 24-18-37.Description of tanks 4 - 500 Barrel wood.Logs of the above wells were filed with the Oil Conservation Commission On well completion 1939All other requirements of the Commission have ~~XXXXXX~~ (have not) been complied with. (Cross out incorrect words.)

Additional information:

Oil from this lease prior to July 13, 1939, run by Shell Pipe Line Corporation.
 Effective July 13, Humble Pipe Line Company assumed the connection.

We request your authorization to the Humble Pipe Line Company to run this oil.



Yours truly,

Permission is hereby granted to make pipe line connections requested above.

OIL CONSERVATION COMMISSION,

By Roy Garbrough
A. ANDREAS
 Title State Geologist
 Date Member Oil Conservation Commission AUG 23 '39

AMERADA PETROLEUM CORPORATION

Owner or Operator

By J. L. Low
 Position Superintendent
 Address Monument, New Mexico.

THE UNIVERSITY OF CHICAGO
DEPARTMENT OF CHEMISTRY

PHYSICAL CHEMISTRY LABORATORY

REPORT ON THE EXPERIMENT DETERMINATION OF THE RATE OF REACTION OF HYDROGEN PEROXIDE WITH POTASSIUM IODIDE

- Name: [Name]
- Date: [Date]

1. Introduction

The purpose of this experiment is to determine the rate of reaction of hydrogen peroxide with potassium iodide.

The reaction is as follows: $2H_2O_2 \rightarrow 2H_2O + O_2$

The rate of reaction is determined by measuring the volume of oxygen gas evolved over a period of time. The rate of reaction is expected to be first order with respect to the concentration of hydrogen peroxide and second order with respect to the concentration of potassium iodide.

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2. Experimental Procedure

The reaction is as follows: $2H_2O_2 \rightarrow 2H_2O + O_2$

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