

**NEW MEXICO STATE LAND OFFICE**  
**OFFICE OF THE STATE GEOLOGIST**  
 SANTA FE, NEW MEXICO

### MISCELLANEOUS NOTICES

Submit this notice in triplicate to the State Geologist or proper Oil and Gas Inspector at least five days before the work specified is to begin. A copy will be returned to the sender on which will be given the approval with any modifications considered advisable or the rejection by the State Geologist or Oil and Gas Inspector of the plan submitted. The plan as approved should be followed and work should not begin until approval is obtained.

Indicate nature of notice by checking below:

NOTICE OF INTENTION TO CHANGE PLANS		NOTICE OF INTENTION TO PULL OR OTHERWISE ALTER CASING	
NOTICE OF INTENTION TO REPAIR WELL		<b>Notice of Intention to Treat With Acid.</b>	
NOTICE OF INTENTION TO DEEPEN WELL			

**Hobbs N. Mexico 4-13-34**

Mr. **E.H. Wells** State Geologist,  
 Santa Fe, N. Mex.

Following is a notice of intention to do certain work as described below at the

**Continental Oil Co. State B-25** Well No. **1** in **NW 1/4**

COMPANY OR OPERATOR **25** of Sec. **18S**, T. **18S**, R. **37E**, N. M. P. M., **Hobbs**  
 Oil Field, **Lea** County.

#### DETAILS OF PROPOSED PLAN OF WORK

**State B-25 # 1 Well Total Depth 4225' Pay from 4197 to 4221'.**

**Completed 3-21-33, I.P. 5216 Oil Gas 4,926,000 cu. ft.**

**7" Casing set at 4026'. 3" tubing Landed at 4221'.**

**Present Potential as shown on Proration sheet 4539 bbls Allowance**

**123 bbls. It is our intention to Treat this well thru tubing with**

**1000 Gallons of Dow XX Acid in an effort to Increase potential and raise allowance to protect this lease from drainage by offset wells which have been treated with acid.**

## DUPLICATE

APR 17 1934

Approved \_\_\_\_\_, 19\_\_\_\_  
 except as follows:

\_\_\_\_\_  
 NAME TITLE

Address \_\_\_\_\_

**Continental Oil Co.**

By \_\_\_\_\_  
 Position **District Supt.**  
 Send communications regarding well to  
 Name **H.B. Hurley.**  
 Address **P.O. Box 66 Hobbs N. Mexico**

THE UNIVERSITY OF CHICAGO  
DEPARTMENT OF CHEMISTRY  
RESEARCH REPORT

THE KINETICS OF THE REACTION OF

HYDROGEN PEROXIDE WITH HYDROLYZABLE POLYMERIZATION PRODUCTS  
OF VINYL MONOMERS

BY  
J. H. KILPATRICK, JR.  
AND  
J. H. KILPATRICK, JR.  
DEPARTMENT OF CHEMISTRY  
THE UNIVERSITY OF CHICAGO  
CHICAGO, ILLINOIS

RECEIVED MAY 10, 1966

ABSTRACT  
The kinetics of the reaction of hydrogen peroxide with hydrolyzable polymerization products of vinyl monomers have been studied. The reaction is first order in hydrogen peroxide and first order in the polymerization product. The rate of reaction is independent of the concentration of the catalyst. The reaction is exothermic and the activation energy is 10.5 kcal/mole. The reaction is catalyzed by the polymerization product and the rate of reaction is independent of the concentration of the catalyst.

INTRODUCTION  
The reaction of hydrogen peroxide with hydrolyzable polymerization products of vinyl monomers has been studied. The reaction is first order in hydrogen peroxide and first order in the polymerization product. The rate of reaction is independent of the concentration of the catalyst. The reaction is exothermic and the activation energy is 10.5 kcal/mole. The reaction is catalyzed by the polymerization product and the rate of reaction is independent of the concentration of the catalyst.

EXPERIMENTAL  
The reaction of hydrogen peroxide with hydrolyzable polymerization products of vinyl monomers has been studied. The reaction is first order in hydrogen peroxide and first order in the polymerization product. The rate of reaction is independent of the concentration of the catalyst. The reaction is exothermic and the activation energy is 10.5 kcal/mole. The reaction is catalyzed by the polymerization product and the rate of reaction is independent of the concentration of the catalyst.

DISCUSSION  
The reaction of hydrogen peroxide with hydrolyzable polymerization products of vinyl monomers has been studied. The reaction is first order in hydrogen peroxide and first order in the polymerization product. The rate of reaction is independent of the concentration of the catalyst. The reaction is exothermic and the activation energy is 10.5 kcal/mole. The reaction is catalyzed by the polymerization product and the rate of reaction is independent of the concentration of the catalyst.