

NEW MEXICO STATE LAND OFFICE
OFFICE OF THE STATE GEOLOGIST
 SANTA FE, NEW MEXICO

MISCELLANEOUS REPORTS ON WELLS

Submit this report in duplicate to the State Geologist or proper Oil and Gas Inspector within ten days after the work specified is completed. It should be signed and sworn to before a notary public for reports on beginning drilling operations, results of shooting well, results of test of water shut-off, result of abandonment of well, and other important operations, even though the work was witnessed by the State Geologist or Oil and Gas Inspector. Reports on minor operations need not be signed and sworn to before a notary public, but such operations should be witnessed by an Oil and Gas Inspector if possible.

Indicate nature of report by checking below:

REPORT ON BEGINNING DRILLING OPERATIONS	X	REPORT ON DEEPENING WELL	
REPORT ON RESULT OF SHOOTING WELL		REPORT ON PULLING OR OTHERWISE ALTERING CASING	
REPORT ON RESULT OF TEST OF WATER SHUT-OFF		REPORT ON REPAIRING WELL	
REPORT ON RESULT OF ABANDONMENT OF WELL			

Hobbs, N.M. 2-3-33

PLACE

DATE

Mr. E.H. Wells State Geologist,

Santa Fe, N. Mex.

Following is a report on the work done and the results obtained under the heading noted above at the

Continental Oil Company State B 25 Well No. 1 in the
 NE 1/4 of NW 1/4 of Sec. 25, T. 18 South, R. 37 East, N. M. P. M.,
Hobbs Oil Field, Lea County.

The dates of this work were as follows: January 29-1933

Notice of intention to do the work was (was not) submitted on Form SG 101 on
Jan. 22-, 1933, and approval of the proposed plan was ~~was not~~ obtained. (Cross
 out incorrect words.)

DETAILED ACCOUNT OF WORK DONE AND RESULTS OBTAINED

There have been no changes in the proposed plan of drilling or casing
 Program, as reported on Form SG 101.

A spudding machine was moved in and drilling commenced on January 29, 1933.

DUPLICATE

Subscribed and sworn to before me this
10th day of Feb, 1933

[Signature]
 NOTARY PUBLIC.
 My commission expires July 2, 1933

Remarks:

I hereby swear or affirm that the information
 given above is true and correct.

Name [Signature]
 Position District Superintendent
 Representing Continental Oil Company
 COMPANY OR OPERATOR.
 Address Box 66 Hobbs, N.M.

APPROVED AS O. K.

[Signature]

THE UNIVERSITY OF CHICAGO
DEPARTMENT OF CHEMISTRY
LABORATORY OF PHYSICAL CHEMISTRY

1954

THEORY OF THE ELECTROLYTIC DECOMPOSITION OF
SOLUBLE ELECTROLYTES

by
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

and
J. H. DINEEN

Abstract: The theory of the electrolytic decomposition of soluble electrolytes is presented. The theory is based on the assumption that the electrolyte is a mixture of a strong electrolyte and a weak electrolyte. The strong electrolyte is assumed to be completely dissociated into ions, and the weak electrolyte is assumed to be in equilibrium with its ions. The theory is applied to the case of the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte. The results show that the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte is a function of the concentration of the strong electrolyte and the concentration of the weak electrolyte. The theory is also applied to the case of the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte. The results show that the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte is a function of the concentration of the strong electrolyte and the concentration of the weak electrolyte.

1. Introduction. The theory of the electrolytic decomposition of soluble electrolytes is presented. The theory is based on the assumption that the electrolyte is a mixture of a strong electrolyte and a weak electrolyte. The strong electrolyte is assumed to be completely dissociated into ions, and the weak electrolyte is assumed to be in equilibrium with its ions. The theory is applied to the case of the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte. The results show that the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte is a function of the concentration of the strong electrolyte and the concentration of the weak electrolyte. The theory is also applied to the case of the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte. The results show that the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte is a function of the concentration of the strong electrolyte and the concentration of the weak electrolyte.

2. Theory. The theory of the electrolytic decomposition of soluble electrolytes is presented. The theory is based on the assumption that the electrolyte is a mixture of a strong electrolyte and a weak electrolyte. The strong electrolyte is assumed to be completely dissociated into ions, and the weak electrolyte is assumed to be in equilibrium with its ions. The theory is applied to the case of the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte. The results show that the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte is a function of the concentration of the strong electrolyte and the concentration of the weak electrolyte. The theory is also applied to the case of the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte. The results show that the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte is a function of the concentration of the strong electrolyte and the concentration of the weak electrolyte.

3. Results. The results of the theory of the electrolytic decomposition of soluble electrolytes are presented. The results show that the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte is a function of the concentration of the strong electrolyte and the concentration of the weak electrolyte. The results are also presented for the case of the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte. The results show that the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte is a function of the concentration of the strong electrolyte and the concentration of the weak electrolyte.

4. Conclusion. The theory of the electrolytic decomposition of soluble electrolytes is presented. The theory is based on the assumption that the electrolyte is a mixture of a strong electrolyte and a weak electrolyte. The strong electrolyte is assumed to be completely dissociated into ions, and the weak electrolyte is assumed to be in equilibrium with its ions. The theory is applied to the case of the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte. The results show that the electrolytic decomposition of a mixture of a strong electrolyte and a weak electrolyte is a function of the concentration of the strong electrolyte and the concentration of the weak electrolyte.