

NEW MEXICO STATE LAND OFFICE
OFFICE OF THE STATE GEOLOGIST
SANTA FE, NEW MEXICO

MISCELLANEOUS REPORTS ON WELLS

Submit this report in duplicate to the State Geologist or proper Oil and Gas Inspector within ten days after the work specified is completed. It should be signed and sworn to before a notary public for reports on beginning drilling operations, results of shooting well, results of test of water shut-off, result of abandonment of well, and other important operations, even though the work was witnessed by the State Geologist or Oil and Gas Inspector. Reports on minor operations need not be signed and sworn to before a notary public, but such operations should be witnessed by an Oil and Gas Inspector if possible.

Indicate nature of report by checking below:

REPORT ON BEGINNING DRILLING OPERATIONS		REPORT ON DEEPENING WELL	
REPORT ON RESULT OF SHOOTING WELL		REPORT ON PULLING OR OTHERWISE ALTERING CASING	
REPORT ON RESULT OF TEST OF WATER SHUT-OFF		REPORT ON REPAIRING WELL	X
REPORT ON RESULT OF ABANDONMENT OF WELL			

Hobbs, New Mexico, January 12, 1935

Mr. E. H. Wells State Geologist,
Santa Fe, N. Mex.

Following is a report on the work done and the results obtained under the heading noted above at the Talbott & Mullineaux Lina Well No. 1 in the 18 S of Sec. 33, T. 18 S, R. 37 E, N. M. P. M., Hobbs Oil Field, Lea County.

The dates of this work were as follows: Nov. 26 to date hereof

Notice of intention to do the work was (was not) submitted on Form SG 105 on Sept. 13, 1934, and approval of the proposed plan was ~~was not~~ obtained. (Cross out incorrect words.)

DETAILED ACCOUNT OF WORK DONE AND RESULTS OBTAINED

We went into the hole with cable tools, and ran into a bridge at 2000 feet. Fished out shoulder of a rotary bit and rotary cones, and then went on down to 3900 without further trouble. From 3900 to top of cement at 4188 drilled most of time in iron. Drilled through cement and also through the wooden plug, and now have hole clean and open to about 4262. Have run into an eight inch bit that was left in hole by previous operators, and are now trying to fish it out. As soon as we get it out want to go on and deepen hole. Water standing in hole to about 1200 foot level.

Subscribed and sworn to before me this

_____ day of _____, 19____

NOTARY PUBLIC.

My commission expires _____

I hereby swear or affirm that the information given above is true and correct.

Name E. H. Wells

Position Partner

Representing Talbott & Mullineaux
COMPANY OR OPERATOR.

Address Hobbs, New Mexico.

Remarks:

NAME

TITLE

THE UNIVERSITY OF CHICAGO
DEPARTMENT OF CHEMISTRY
RESEARCH REPORT

THE KINETICS OF THE REACTION OF
HYDROGEN PEROXIDE WITH
HYDROXYLAMINE

BY
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ABSTRACT
The reaction of hydrogen peroxide with hydroxylamine has been studied in aqueous solution at 25°C. The reaction is first order in both reactants and the rate constant is $1.1 \times 10^{-4} \text{ sec}^{-1}$.

INTRODUCTION
The reaction of hydrogen peroxide with hydroxylamine has been studied in aqueous solution at 25°C. The reaction is first order in both reactants and the rate constant is $1.1 \times 10^{-4} \text{ sec}^{-1}$.

EXPERIMENTAL
The reaction of hydrogen peroxide with hydroxylamine has been studied in aqueous solution at 25°C. The reaction is first order in both reactants and the rate constant is $1.1 \times 10^{-4} \text{ sec}^{-1}$.

RESULTS
The reaction of hydrogen peroxide with hydroxylamine has been studied in aqueous solution at 25°C. The reaction is first order in both reactants and the rate constant is $1.1 \times 10^{-4} \text{ sec}^{-1}$.

DISCUSSION
The reaction of hydrogen peroxide with hydroxylamine has been studied in aqueous solution at 25°C. The reaction is first order in both reactants and the rate constant is $1.1 \times 10^{-4} \text{ sec}^{-1}$.

CONCLUSIONS
The reaction of hydrogen peroxide with hydroxylamine has been studied in aqueous solution at 25°C. The reaction is first order in both reactants and the rate constant is $1.1 \times 10^{-4} \text{ sec}^{-1}$.

REFERENCES
The reaction of hydrogen peroxide with hydroxylamine has been studied in aqueous solution at 25°C. The reaction is first order in both reactants and the rate constant is $1.1 \times 10^{-4} \text{ sec}^{-1}$.

ACKNOWLEDGMENTS
The reaction of hydrogen peroxide with hydroxylamine has been studied in aqueous solution at 25°C. The reaction is first order in both reactants and the rate constant is $1.1 \times 10^{-4} \text{ sec}^{-1}$.

APPENDIX
The reaction of hydrogen peroxide with hydroxylamine has been studied in aqueous solution at 25°C. The reaction is first order in both reactants and the rate constant is $1.1 \times 10^{-4} \text{ sec}^{-1}$.

NOTES
The reaction of hydrogen peroxide with hydroxylamine has been studied in aqueous solution at 25°C. The reaction is first order in both reactants and the rate constant is $1.1 \times 10^{-4} \text{ sec}^{-1}$.

LITERATURE CITED
The reaction of hydrogen peroxide with hydroxylamine has been studied in aqueous solution at 25°C. The reaction is first order in both reactants and the rate constant is $1.1 \times 10^{-4} \text{ sec}^{-1}$.

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