

NEW MEXICO STATE LAND OFFICE  
OFFICE OF THE STATE GEOLOGIST  
SANTA FE, NEW MEXICO

### MISCELLANEOUS NOTICES

Submit this notice in triplicate to the State Geologist or proper Oil and Gas Inspector at least five days before the work specified is to begin. A copy will be returned to the sender on which will be given the approval with any modifications considered advisable or the rejection by the State Geologist or Oil and Gas Inspector of the plan submitted. The plan as approved should be followed and work should not begin until approval is obtained.

Indicate nature of notice by checking below:

NOTICE OF INTENTION TO CHANGE PLANS		NOTICE OF INTENTION TO PULL OR OTHERWISE ALTER CASING	
NOTICE OF INTENTION TO REPAIR WELL		<b>ACID TREAT WELL</b>	<b>X</b>
NOTICE OF INTENTION TO DEEPEN WELL			

Hobbs, New Mexico, August 25, 1934

Mr. J. D. Hunter, ~~State Geologist~~

~~State Oil & Gas Inspector~~ State Oil & Gas Inspector, Carlsbad, New Mexico.

Following is a notice of intention to do certain work as described below at the Shell

Petroleum Corporation-W.D.Grimes

Well No. 4 in S.W.1/4

COMPANY OR OPERATOR

of Sec. 28, T. 18S, R. 38E, N. M. P. M., Hobbs

Oil Field, Lea County.

#### DETAILS OF PROPOSED PLAN OF WORK

**Well will be treated with 1000 gal. 50% acid by Chemical Process Company from 3935 to 4090 ft.**

DUPLICATE

OCT 2 1934

Approved \_\_\_\_\_, 19\_\_\_\_  
except as follows:

J. D. Hunter  
NAME by F.H. TITLE \_\_\_\_\_

Address \_\_\_\_\_

**SHELL PETROLEUM CORPORATION**

By R. J. Schell COMPANY OR OPERATOR

Position District Engineer

Send communications regarding well to

Name Shell Petroleum Corporation

Address Box 996, Wink, Texas.

ICR.

THE UNIVERSITY OF CHICAGO  
DEPARTMENT OF CHEMISTRY  
RESEARCH REPORT

PHOTOCHEMICAL REACTIONS

The following report describes the results of a study of the photochemical reactions of certain organic compounds. The study was carried out in the Department of Chemistry, University of Chicago, under the direction of Professor [Name].

The compounds studied were [Name] and [Name]. The reactions were carried out in [Name] solution, under [Name] conditions.

The results of the study are as follows: [Name] reacts with [Name] to form [Name]. The reaction is first order in [Name] and second order in [Name]. The rate constant for the reaction is [Name].

The following table gives the rate constants for the reaction of [Name] with [Name] at various temperatures:

[Table with 2 columns: Temperature (°C) and Rate Constant (l./mole-sec.)]

The results show that the rate constant for the reaction of [Name] with [Name] increases with increasing temperature. This is in agreement with the Arrhenius equation.

The following table gives the activation energies for the reaction of [Name] with [Name] at various temperatures:

[Table with 2 columns: Temperature (°C) and Activation Energy (kcal/mole)]

The results show that the activation energy for the reaction of [Name] with [Name] is [Name] kcal/mole. This is in agreement with the Arrhenius equation.

The following table gives the quantum yields for the reaction of [Name] with [Name] at various wavelengths:

[Table with 2 columns: Wavelength (nm) and Quantum Yield]

The results show that the quantum yield for the reaction of [Name] with [Name] is [Name]. This is in agreement with the Arrhenius equation.

The following table gives the rate constants for the reaction of [Name] with [Name] at various wavelengths:

[Table with 2 columns: Wavelength (nm) and Rate Constant (l./mole-sec.)]

The results show that the rate constant for the reaction of [Name] with [Name] increases with increasing wavelength. This is in agreement with the Arrhenius equation.

The following table gives the activation energies for the reaction of [Name] with [Name] at various wavelengths:

[Table with 2 columns: Wavelength (nm) and Activation Energy (kcal/mole)]

The results show that the activation energy for the reaction of [Name] with [Name] is [Name] kcal/mole. This is in agreement with the Arrhenius equation.

The following table gives the quantum yields for the reaction of [Name] with [Name] at various wavelengths:

[Table with 2 columns: Wavelength (nm) and Quantum Yield]