

Indian Agency.....

Allottee.....

Lease No. **Louis C. Popeano**

Account #1

(SUBMIT IN TRIPLICATE)

UNITED STATES
DEPARTMENT OF THE INTERIOR
GEOLOGICAL SURVEY

Las Cruces
U. S. Land Office.....

045741

Lease or permit No.

DUPLICATE

SUNDRY NOTICES AND REPORTS ON WELLS

NOTICE OF INTENTION TO DRILL.....	SUBSEQUENT REPORT OF WATER SHUT-OFF.....
NOTICE OF INTENTION TO CHANGE PLANS.....	SUBSEQUENT REPORT OF SHOOTING OR ACIDIZING.....
NOTICE OF INTENTION TO TEST WATER SHUT-OFF.....	SUBSEQUENT REPORT OF ALTERING CASING.....
NOTICE OF INTENTION TO RE-DRILL OR REPAIR WELL.....	SUBSEQUENT REPORT OF REDRILLING OR REPAIR.....
NOTICE OF INTENTION TO SHOOT OR ACIDIZE.....	SUBSEQUENT REPORT OF ABANDONMENT.....
NOTICE OF INTENTION TO PULL OR ALTER CASING.....	SUPPLEMENTARY WELL HISTORY.....
NOTICE OF INTENTION TO ABANDON WELL.....	

(INDICATE ABOVE BY CHECK MARK NATURE OF REPORT, NOTICE, OR OTHER DATA)

Midland, Texas, August 11, 1938

Well No. **6** is located **660** ft. from **[N]** line and **660** ft. from **[E]** line of sec. **35**
NE/4 of NE/4 of Sec. 35 **20-South** **36-East**
(1/4 Sec. and Sec. No.) (Twp.) (Range) (Meridian)
Eunice **Lea** **New Mexico**
(Field) (County or Subdivision) (State or Territory)

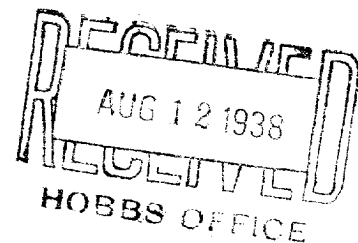
The elevation of the derrick floor above sea level is **3566** ft.

DETAILS OF WORK

(State names of and expected depths to objective sands; show sizes, weights, and lengths of proposed casings; indicate mudding jobs, cementing points, and all other important proposed work)

Tested 10-3/4" casing by bailing dry & drilled plug on August 10, 1938 -
tested O. K.

No change in methods previously reported on Form 9-331a dated August 10, 1938.



I understand that this plan of work must receive approval in writing by the Geological Survey before operations may be commenced.

Company **Humble Oil & Refining Company**

Address **Box 1600 - Midland, Texas.**

By *W. S. Shultz*
Title **Division Superintendent**

THE UNIVERSITY OF CHICAGO
DEPARTMENT OF CHEMISTRY
RESEARCH REPORT NO. 1000
1955

THE EFFECT OF TEMPERATURE ON THE
RATE OF REACTION OF HYDROGEN
PEROXIDE WITH FERROUS SULFATE

BY
J. H. KILPATRICK
AND
J. E. HARRIS

DEPARTMENT OF CHEMISTRY
UNIVERSITY OF CHICAGO
CHICAGO, ILLINOIS

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ABSTRACT
The rate of reaction of hydrogen peroxide with ferrous sulfate has been studied as a function of temperature, concentration of reactants, and the presence of various catalysts.

INTRODUCTION
The reaction of hydrogen peroxide with ferrous sulfate is a well-known reaction which has been studied extensively in the literature. The reaction is exothermic and is catalyzed by a number of substances.

EXPERIMENTAL
The reaction was studied by measuring the rate of evolution of oxygen gas. The reaction mixture was prepared by adding a known volume of a known concentration of hydrogen peroxide to a known volume of a known concentration of ferrous sulfate.

RESULTS
The rate of reaction increases with increasing temperature and increasing concentration of reactants. The reaction is catalyzed by a number of substances, including copper(II) ions, cerium(IV) ions, and manganese(II) ions.

DISCUSSION
The reaction of hydrogen peroxide with ferrous sulfate is a complex reaction which involves a number of steps. The rate of reaction is determined by the rate of the slowest step in the reaction mechanism.

CONCLUSIONS
The rate of reaction of hydrogen peroxide with ferrous sulfate is a function of temperature, concentration of reactants, and the presence of various catalysts. The reaction is exothermic and is catalyzed by a number of substances.

REFERENCES
1. J. H. Kilpatrick and J. E. Harris, *J. Chem. Phys.*, **24**, 1000 (1956).

2. J. H. Kilpatrick and J. E. Harris, *J. Chem. Phys.*, **24**, 1000 (1956).

3. J. H. Kilpatrick and J. E. Harris, *J. Chem. Phys.*, **24**, 1000 (1956).

4. J. H. Kilpatrick and J. E. Harris, *J. Chem. Phys.*, **24**, 1000 (1956).

5. J. H. Kilpatrick and J. E. Harris, *J. Chem. Phys.*, **24**, 1000 (1956).