

NEW CO C CONSERVATION COMMISSION
Santa Fe, New Mexico

MISCELLANEOUS REPORTS ON WELLS

Submit this report in triplicate to the Oil Conservation Commission or its proper agent within ten days after the work specified is completed. It should be signed and sworn to before a notary public for reports on beginning drilling operations, results of shooting well, results of test of casing shut-off, result of plugging of well, and other important operations, even though the work was witnessed by an agent of the Commission. Reports on minor operations need not be signed and sworn to before a notary public. See additional instructions in the Rules and Regulations of the Commission.

Indicate nature of report by checking below:

REPORT ON BEGINNING DRILLING OPERATIONS		REPORT ON REPAIRING WELL	
REPORT ON RESULT OF SHOOTING OR CHEMICAL TREATMENT OF WELL		REPORT ON PULLING OR OTHERWISE ALTERING CASING	
REPORT ON RESULT OF TEST OF CASING SHUT-OFF	X	REPORT ON DEEPENING WELL	
REPORT ON RESULT OF PLUGGING OF WELL			

Midland, Texas

August 28, 1936

Place

Date

OIL CONSERVATION COMMISSION,
Santa Fe, New Mexico.

Gentlemen:

Following is a report on the work done and the results obtained under the heading noted above at the _____

Humble Oil & Refining Company M. E. Langhlin Well No. 1 in the
Company or Operator Lease
SE 1/4 of NE 1/4 of Sec. 4 T. 20-S R. 37-E N. M. P. M.,
Monument Field, Lee County.

The dates of this work were as follows: August 24, 1936

Notice of intention to do the work was [~~submitted~~] submitted on Form C-102 on August 25 1936
and approval of the proposed plan was [~~received~~] obtained. (Cross out incorrect words.)

DETAILED ACCOUNT OF WORK DONE AND RESULTS OBTAINED

Tested 5-1/8" casing with 1200# cold water pressure and drilled plug at 6:00 A. M.

August 27, 1936.

Tested O. K.

DUPLICATE

Witnessed by H. A. Sagie Texas Company Pumper
Name Company Title

Subscribed and sworn to before me this _____

28th day of August, 19 36

Hate Wilson Alard
Notary Public

My Commission expires 6-1-37

I hereby swear or affirm that the information given above is true and correct.

Name [Signature]

Position Division Superintendent

Representing Humble Oil & Refining Company
Company or Operator

Address Drawer W, Midland, Texas

Remarks:

APPROVED

[Signature] Name

[Signature] Title

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2000 1999 1998 1997 1996 1995 1994 1993 1992 1991 1990 1989 1988 1987 1986 1985 1984 1983 1982 1981 1980 1979 1978 1977 1976 1975 1974 1973 1972 1971 1970 1969 1968 1967 1966 1965 1964 1963 1962 1961 1960 1959 1958 1957 1956 1955 1954 1953 1952 1951 1950 1949 1948 1947 1946 1945 1944 1943 1942 1941 1940 1939 1938 1937 1936 1935 1934 1933 1932 1931 1930 1929 1928 1927 1926 1925 1924 1923 1922 1921 1920 1919 1918 1917 1916 1915 1914 1913 1912 1911 1910 1909 1908 1907 1906 1905 1904 1903 1902 1901 1900

[illegible]

DOI: 10.1002/for

[illegible]

1. What is the purpose of the document?

Figure 1. The effect of the concentration of the H_2O_2 solution on the amount of the H_2O_2 consumed in the reaction of the H_2O_2 with the Fe^{2+} ion. The concentration of the Fe^{2+} ion was 1.0×10^{-3} mol/L, and the concentration of the H_2O_2 solution was 0.001, 0.002, 0.004, 0.006, 0.008, 0.010, 0.012, 0.014, 0.016, 0.018, 0.020, 0.022, 0.024, 0.026, 0.028, 0.030, 0.032, 0.034, 0.036, 0.038, 0.040, 0.042, 0.044, 0.046, 0.048, 0.050, 0.052, 0.054, 0.056, 0.058, 0.060, 0.062, 0.064, 0.066, 0.068, 0.070, 0.072, 0.074, 0.076, 0.078, 0.080, 0.082, 0.084, 0.086, 0.088, 0.090, 0.092, 0.094, 0.096, 0.098, 0.100 mol/L. The concentration of the Fe^{2+} ion was 1.0×10^{-3} mol/L, and the concentration of the H_2O_2 solution was 0.001, 0.002, 0.004, 0.006, 0.008, 0.010, 0.012, 0.014, 0.016, 0.018, 0.020, 0.022, 0.024, 0.026, 0.028, 0.030, 0.032, 0.034, 0.036, 0.038, 0.040, 0.042, 0.044, 0.046, 0.048, 0.050, 0.052, 0.054, 0.056, 0.058, 0.060, 0.062, 0.064, 0.066, 0.068, 0.070, 0.072, 0.074, 0.076, 0.078, 0.080, 0.082, 0.084, 0.086, 0.088, 0.090, 0.092, 0.094, 0.096, 0.098, 0.100 mol/L. The concentration of the Fe^{2+} ion was 1.0×10^{-3} mol/L, and the concentration of the H_2O_2 solution was 0.001, 0.002, 0.004, 0.006, 0.008, 0.010, 0.012, 0.014, 0.016, 0.018, 0.020, 0.022, 0.024, 0.026, 0.028, 0.030, 0.032, 0.034, 0.036, 0.038, 0.040, 0.042, 0.044, 0.046, 0.048, 0.050, 0.052, 0.054, 0.056, 0.058, 0.060, 0.062, 0.064, 0.066, 0.068, 0.070, 0.072, 0.074, 0.076, 0.078, 0.080, 0.082, 0.084, 0.086, 0.088, 0.090, 0.092, 0.094, 0.096, 0.098, 0.100 mol/L.

$\frac{d}{dt} \left(\frac{\partial L}{\partial \dot{x}} \right) = \frac{\partial L}{\partial x}$

Journal of Management Studies, 20(6), 791-806.

[illegible]

ALL INFORMATION CONTAINED HEREIN IS UNCLASSIFIED

1. The following information is being furnished to you for your information only. It is not to be used for any other purpose.

• *Journal of the American Medical Association*, 1997; 277: 1033-1037

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100-443887-1000

[illegible]

100-441100-1000

1. The first step is to identify the problem or question that needs to be answered. This involves understanding the context and the specific requirements of the task.

$\frac{d}{dt} \left(\frac{1}{\rho} \right) = - \frac{1}{\rho^2} \frac{d\rho}{dt}$