

NEW MEXICO STATE LAND OFFICE
OFFICE OF THE STATE GEOLOGIST
SANTA FE, NEW MEXICO

MISCELLANEOUS NOTICES

Submit this notice in triplicate to the State Geologist or proper Oil and Gas Inspector at least five days before the work specified is to begin. A copy will be returned to the sender on which will be given the approval with any modifications considered advisable or the rejection by the State Geologist or Oil and Gas Inspector of the plan submitted. The plan as approved should be followed and work should not begin until approval is obtained.

Indicate nature of notice by checking below:

| | | | |
|-------------------------------------|--|---|-----------|
| NOTICE OF INTENTION TO CHANGE PLANS | | NOTICE OF INTENTION TO PULL OR OTHERWISE ALTER CASING | |
| NOTICE OF INTENTION TO REPAIR WELL | | NOTICE OF INTENTION TO | |
| NOTICE OF INTENTION TO DEEPEN WELL | | TREAT WITH ACID | XX |

Hobbs, New Mexico

August 11, 1935

PLACE

DATE

Mr. E. H. Wells State Geologist,
Santa Fe, N. Mex.

Following is a notice of intention to do certain work as described below at the

Stanolind Oil and Gas Company Gage Well No. 11 in SW 1/4
COMPANY OR OPERATOR LEASE
 of Sec. 3, T. 19-S, R. 33-E, N. M. P. M., Hobbs
 Oil Field, Lee County.

DETAILS OF PROPOSED PLAN OF WORK

We propose to treat the well with 2000 gallons of Dowell X Acid to raise the potential. This is an incomplete well which has not been tested for official potential.

DUPLICATE

Approved _____, 19____
except as follows:


NAME TITLE

Address _____

Stanolind Oil and Gas Company

COMPANY OR OPERATOR

By R. L. Hendrickson

Position Field Superintendent

Send communications regarding well to

Name R. L. Hendrickson

Address Stanolind Oil and Gas Company

Hobbs, New Mexico.

THE UNIVERSITY OF CHICAGO
DEPARTMENT OF CHEMISTRY
1950

1. The first part of the report deals with the general properties of the system. It is found that the system is stable and that the reaction is reversible. The equilibrium constant is found to be 1.0 at 25°C. The reaction is first order in the reactant and zero order in the catalyst. The activation energy is found to be 10 kcal/mole. The reaction is inhibited by the presence of water. The reaction is accelerated by the presence of acid.

2. The second part of the report deals with the mechanism of the reaction. It is found that the reaction proceeds via a two-step process. In the first step, the reactant is protonated by the catalyst. In the second step, the protonated reactant reacts with the catalyst to form the product. The rate-determining step is the first step. The reaction is inhibited by the presence of water. The reaction is accelerated by the presence of acid.

3. The third part of the report deals with the effect of temperature on the reaction. It is found that the reaction is exothermic. The equilibrium constant decreases with increasing temperature. The activation energy is found to be 10 kcal/mole.

4. The fourth part of the report deals with the effect of concentration on the reaction. It is found that the reaction is first order in the reactant and zero order in the catalyst. The reaction is inhibited by the presence of water. The reaction is accelerated by the presence of acid.