

OIL CONSERVATION COMMISSION

BOX 2045

HOBBS, NEW MEXICO

DATE November 4, 1960

OIL CONSERVATION COMMISSION
BOX 871
SANTA FE, NEW MEXICO

Re: Proposed NSP _____

Proposed NSL _____

Proposed NFC _____

Proposed DC X

Gentlemen:

I have examined the application dated _____
for the Continental Britt B-15 #11 N 15-20-37
Operator Lease and Well No. S-T-R

and my recommendations are as follows:

I recommend that this D. C. be approved due to the corrected filing
of new perforations, packer, and formation tops.* JWR

OK-EFE

Yours very truly,

OIL CONSERVATION COMMISSION

Engineer District I

* Log Tops: Tubb 6324', Drinkard 6647'
Tubb Perf. 6540 to 6600'
Drinkard Perf. 6817-48'
Packer About 6640'

THEORY

The theory of the present experiment is based on the fact that the rate of reaction between a substance and a reagent is proportional to the concentration of the substance. In this case, the substance is the unknown concentration of the solution, and the reagent is the standard solution. The rate of reaction is measured by the time taken for the reaction to complete, which is inversely proportional to the concentration of the substance. The standard solution is a known concentration of the same substance, and its rate of reaction is used as a reference to determine the concentration of the unknown solution.

The standard solution is prepared by weighing a precise amount of the substance and dissolving it in a known volume of solvent. The unknown solution is prepared by weighing a precise amount of the substance and dissolving it in a known volume of solvent. The rate of reaction for the standard solution is measured by the time taken for the reaction to complete. The rate of reaction for the unknown solution is measured by the time taken for the reaction to complete. The concentration of the unknown solution is determined by comparing the rate of reaction of the unknown solution to the rate of reaction of the standard solution.

The concentration of the unknown solution is determined by the following equation:

$$C_{\text{unknown}} = \frac{t_{\text{standard}}}{t_{\text{unknown}}} \times C_{\text{standard}}$$

where C_{unknown} is the concentration of the unknown solution, t_{standard} is the time taken for the reaction to complete for the standard solution, t_{unknown} is the time taken for the reaction to complete for the unknown solution, and C_{standard} is the concentration of the standard solution.

EXPERIMENTAL PROCEDURE

1. Preparation of Standard Solution

1. Weigh a precise amount of the substance (e.g., 0.1000 g) and transfer it to a volumetric flask.

2. Dissolve the substance in a small amount of solvent (e.g., water) and then dilute to the mark with solvent.

3. The concentration of the standard solution is determined by the following equation:

$$C_{\text{standard}} = \frac{m}{V}$$

where C_{standard} is the concentration of the standard solution, m is the mass of the substance, and V is the volume of the solution.