

## NEW MEXICO OIL CONSERVATION COMMISSION

Santa Fe, New Mexico

## MISCELLANEOUS REPORTS ON WELL

Submit this report in triplicate to the Oil Conservation Commission or its proper agent within ten days after the work specified is completed. It should be signed and sworn to before a notary public for reports on beginning drilling operations, results of shooting well, results of test of casing shut-offs, result of plugging of well, and other important operations, even though the work was witnessed by an agent of the commission. Reports on minor operations need not be signed and sworn to before a notary public. See additional instructions in the Rules and Regulations of the Commission.

Indicate nature of report by checking below:

REPORT ON BEGINNING DRILLING OPERATIONS		REPORT ON REPAIRING WELL	
REPORT ON RESULT OF SHOOTING OR CHEMICAL TREATMENT OF WELL		REPORT ON PULLING OR OTHERWISE ALTERING CASING	
REPORT ON RESULT OF TEST OF CASING SHUT-OFF		REPORT ON DEEPENING WELL	
REPORT ON RESULT OF PLUGGING OF WELL		Acidizing	xx

Hobbs, New Mexico

November 25, 1939

Place

Date

OIL CONSERVATION COMMISSION,  
Santa Fe, New Mexico.

Gentlemen:

Following is a report on the work done and the results obtained under the heading noted above at the

Gulf Oil Corpn. - Gypsy Div.

Mark Owen

Well No. 1

in the

Company or Operator

Lease

NW SE

of Sec. 34

T. 21

R. 37

N. M. P. M.,

Hardy

Field,

Lea

County

The dates of this work were as follows: Acidized November 25, 1939

Notice of intention to do the work was (was not) submitted on Form C-102 on \_\_\_\_\_ 19\_\_\_\_  
and approval of the proposed plan was (was not) obtained. (Cross out incorrect words)

## DETAILED ACCOUNT OF WORK DONE AND RESULTS OBTAINED

Acidized with 2,000 gals, 1000 Gals Destablized and 1,000 gals. 15%, By Halliburton.

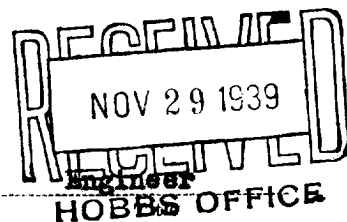
Production before acid:- No test.

Production after test:- Flowed 75 barrels oil in 24 hours, thru 1/2" choke 3 rounds  
open cutting .6% BSW.

Witnessed by R. L. Blynn  
Name

Gulf

Company



Subscribed and sworn to before me this

I hereby swear or affirm that the information given above is true and correct.

25

day of November

1939

Name

Position District Supt.

Representing Gulf Oil Corpn. - Gypsy Prod. Div.  
Company or Operator

My Commission expires February 25, 1942

Address Hobbs, New Mexico

Remarks:

Roy Yarbrough  
Name  
OIL & GAS INSPECTOR  
Title

THE UNIVERSITY OF CHICAGO  
DEPARTMENT OF CHEMISTRY  
RESEARCH REPORT NO. 1000

THE KINETICS OF THE REACTION OF

HYDROGEN PEROXIDE WITH HYDROGEN SULFIDE IN AQUEOUS SOLUTION  
AT 25°C. AND 1 ATMOSPHERE

BY J. H. KINZIE AND J. H. KINZIE

RECEIVED JANUARY 10, 1957  
REVISION RECEIVED MARCH 15, 1957  
ACCEPTED FOR PUBLICATION MARCH 22, 1957

ABSTRACT: The reaction of hydrogen peroxide with hydrogen sulfide in aqueous solution at 25°C. and 1 atmosphere has been studied. The reaction is first order in hydrogen peroxide and second order in hydrogen sulfide.

The reaction of hydrogen peroxide with hydrogen sulfide in aqueous solution at 25°C. and 1 atmosphere has been studied. The reaction is first order in hydrogen peroxide and second order in hydrogen sulfide. The rate constant for the reaction is  $1.1 \times 10^{-4} \text{ liter}^2 \text{ mole}^{-2} \text{ sec}^{-1}$ . The activation energy for the reaction is 14.5 kcal/mole. The reaction is catalyzed by iron(II) ions and copper(II) ions. The rate of reaction is independent of the concentration of the catalyst.

The reaction of hydrogen peroxide with hydrogen sulfide in aqueous solution at 25°C. and 1 atmosphere has been studied. The reaction is first order in hydrogen peroxide and second order in hydrogen sulfide. The rate constant for the reaction is  $1.1 \times 10^{-4} \text{ liter}^2 \text{ mole}^{-2} \text{ sec}^{-1}$ . The activation energy for the reaction is 14.5 kcal/mole. The reaction is catalyzed by iron(II) ions and copper(II) ions. The rate of reaction is independent of the concentration of the catalyst.

INTRODUCTION

The reaction of hydrogen peroxide with hydrogen sulfide in aqueous solution at 25°C. and 1 atmosphere has been studied. The reaction is first order in hydrogen peroxide and second order in hydrogen sulfide.

EXPERIMENTAL

The reaction of hydrogen peroxide with hydrogen sulfide in aqueous solution at 25°C. and 1 atmosphere has been studied. The reaction is first order in hydrogen peroxide and second order in hydrogen sulfide. The rate constant for the reaction is  $1.1 \times 10^{-4} \text{ liter}^2 \text{ mole}^{-2} \text{ sec}^{-1}$ . The activation energy for the reaction is 14.5 kcal/mole. The reaction is catalyzed by iron(II) ions and copper(II) ions. The rate of reaction is independent of the concentration of the catalyst.

DISCUSSION

CONCLUSION