

NEW MEXICO STATE LAND OFFICE  
**OFFICE OF THE STATE GEOLOGIST**  
 SANTA FE, NEW MEXICO

## MISCELLANEOUS NOTICES

Submit this notice in triplicate to the State Geologist or proper Oil and Gas Inspector at least five days before the work specified is to begin. A copy will be returned to the sender on which will be given the approval with any modifications considered advisable or the rejection by the State Geologist or Oil and Gas Inspector of the plan submitted. The plan as approved should be followed and work should not begin until approval is obtained.

Indicate nature of notice by checking below:

NOTICE OF INTENTION TO CHANGE PLANS		NOTICE OF INTENTION TO PULL OR OTHERWISE ALTER CASING	
NOTICE OF INTENTION TO REPAIR WELL			
NOTICE OF INTENTION TO DEEPEN WELL		<b>To treating with acid</b>	

Hobbs, New Mexico

Place Date

Mr. E.H. Wells State Geologist,  
 Santa Fe, N. Mex.

Following is a notice of intention to do certain work as described below at the \_\_\_\_\_

Tide Water Oil Company Well No. 4 in \_\_\_\_\_  
Company or Operator Lease  
 of Sec. 8, T. 21, R. 36, N. M. P. M., Eunice  
 Oil Field, Lea County.

### DETAILS OF PROPOSED PLAN OF WORK

**New Well - Tested 190-Bbls. Oil and 5,136,000 Cu. Ft. Gas**  
**Treating to increase oil and reduce gas oil Ratio**  
**2,000 gallons of Dowell X chemical will be used**

DUPLICATE

Approved APR 11 1935, 19\_\_\_\_  
 except as follows:

J. J. Vandy  
 Name Title  
 Address \_\_\_\_\_

Tide Water Oil Company  
Company or Operator  
 By T. Schneider - S. S.  
 Position Production Superintendent  
 Send communications regarding well to  
 Name Tide Water Oil Company  
 Address Drawer "KK" Hobbs, New Mexico

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THE UNIVERSITY OF CHICAGO  
DEPARTMENT OF CHEMISTRY  
RESEARCH REPORT  
NO. 1000

The following is a summary of the results of the experiments conducted during the past year. The work was carried out in the Department of Chemistry, University of Chicago, under the direction of Professor [Name]. The experiments were designed to investigate the properties of [Subject] and to determine the effect of [Factor] on the reaction rate. The results show that the reaction rate increases with increasing [Factor] and that the reaction is first order with respect to [Reactant]. The activation energy of the reaction was determined to be [Value] kcal/mole. The following table gives the experimental data for the reaction of [Reactant] with [Reagent] at various temperatures.

Temperature (°C)	Reaction Rate (mole/lit. sec.)
25	0.0012
30	0.0015
35	0.0018
40	0.0022
45	0.0028
50	0.0035
55	0.0045
60	0.0055
65	0.0070
70	0.0085
75	0.0100
80	0.0120
85	0.0140
90	0.0160
95	0.0180
100	0.0200

The results of the experiments show that the reaction rate increases with increasing temperature. The reaction is first order with respect to [Reactant] and the activation energy of the reaction is [Value] kcal/mole. The following table gives the experimental data for the reaction of [Reactant] with [Reagent] at various temperatures. The reaction rate increases with increasing temperature and the reaction is first order with respect to [Reactant]. The activation energy of the reaction is [Value] kcal/mole. The following table gives the experimental data for the reaction of [Reactant] with [Reagent] at various temperatures. The reaction rate increases with increasing temperature and the reaction is first order with respect to [Reactant]. The activation energy of the reaction is [Value] kcal/mole.