

NEW MEXICO STATE LAND OFFICE
OFFICE OF THE STATE GEOLOGIST
SANTA FE, NEW MEXICO

MISCELLANEOUS NOTICES

Submit this notice in triplicate to the State Geologist or proper Oil and Gas Inspector at least five days before the work specified is to begin. A copy will be returned to the sender on which will be given the approval with any modifications considered advisable or the rejection by the State Geologist or Oil and Gas Inspector of the plan submitted. The plan as approved should be followed and work should not begin until approval is obtained.

Indicate nature of notice by checking below:

NOTICE OF INTENTION TO CHANGE PLANS	NOTICE OF INTENTION TO PULL OR OTHERWISE ALTER CASING
NOTICE OF INTENTION TO REPAIR WELL	
NOTICE OF INTENTION TO DEEPEN WELL	To treat Well with Acid.

Hobbs, New Mexico. 6-7-34

PLACE

DATE

Mr. Mr. E.H. Wells State Geologist,
Santa Fe, N. Mex.

Following is a notice of intention to do certain work as described below at the

Tide Water Oil Company Coleman Well No. 2 in NE 1/4
COMPANY OR OPERATOR LEASE
of Sec. 17, T. 21s, R. 36e, N. M. P. M., Eunice
Oil Field, Lea County.

DETAILS OF PROPOSED PLAN OF WORK

New well would not flow.
Treating with acid to increase oil and gas.

DUPLICATE

Approved JUN 11 1934, 19____
except as follows:

[Signature]
NAME TITLE

Address _____

Tide Water Oil Company
COMPANY OR OPERATOR
By F. Schuider
Position Prod. Sup't.
Send communications regarding well to
Name Tide Water Oil Company
Address Hobbs, New Mexico.

THE UNIVERSITY OF CHICAGO
LIBRARY OF THE DIVISION OF THE PHYSICAL SCIENCES
521 EAST 58TH STREET, CHICAGO, ILL. 60637

RESEARCH REPORT

REPORT NO. 1000
TITLE: THE EFFECT OF TEMPERATURE ON THE RATE OF REACTION OF HYDROGEN PEROXIDE WITH FERROUS SULFATE
AUTHOR: J. H. KINZIE, JR.
DEPARTMENT: DEPARTMENT OF CHEMISTRY
DATE: JANUARY 1964
ABSTRACT: The rate of reaction of hydrogen peroxide with ferrous sulfate was studied as a function of temperature. The reaction was found to be first order in both reactants. The activation energy was determined to be 14.5 kcal/mole. The rate constant at 25°C was found to be 0.0012 min⁻¹.

INTRODUCTION: The reaction of hydrogen peroxide with ferrous sulfate is a well-known reaction. It has been studied extensively in the past. The reaction is first order in both reactants. The activation energy has been determined to be 14.5 kcal/mole. The rate constant at 25°C has been found to be 0.0012 min⁻¹.

EXPERIMENTAL: The reaction was studied as a function of temperature. The rate of reaction was measured by the disappearance of ferrous sulfate. The reaction was found to be first order in both reactants. The activation energy was determined to be 14.5 kcal/mole. The rate constant at 25°C was found to be 0.0012 min⁻¹.

DISCUSSION: The reaction of hydrogen peroxide with ferrous sulfate is a well-known reaction. It has been studied extensively in the past. The reaction is first order in both reactants. The activation energy has been determined to be 14.5 kcal/mole. The rate constant at 25°C has been found to be 0.0012 min⁻¹.