

NEW MEXICO OIL CONSERVATION COMMISSION  
Santa Fe, New Mexico

MISCELLANEOUS REPORTS ON WELLS

Submit this report in triplicate to the Oil Conservation Commission or its proper agent within ten days after the work specified is completed. It should be signed and sworn to before a notary public for reports on beginning drilling operations, results of shooting well, results of test of casing shut-off, result of plugging of well, and other important operations, even though the work was witnessed by an agent of the Commission. Reports on minor operations need not be signed and sworn to before a notary public. See additional instructions in the Rules and Regulations of the Commission.

Indicate nature of report by checking below:

REPORT ON BEGINNING DRILLING OPERATIONS	<input checked="" type="checkbox"/>	REPORT ON REPAIRING WELL	<input type="checkbox"/>
REPORT ON RESULT OF SHOOTING OR CHEMICAL TREATMENT OF WELL	<input type="checkbox"/>	REPORT ON PULLING OR OTHERWISE ALTERING CASING	<input type="checkbox"/>
REPORT ON RESULT OF TEST OF CASING SHUT-OFF	<input type="checkbox"/>	REPORT ON DEEPENING WELL	<input type="checkbox"/>
REPORT ON RESULT OF PLUGGING OF WELL	<input type="checkbox"/>		

Hobbs N.M.

Dec 17 1935

OIL CONSERVATION COMMISSION,  
Santa Fe, New Mexico.

Gentlemen:

Following is a report on the work done and the results obtained under the heading noted above at the \_\_\_\_\_

Shell Petroleum Corporation Devonian State Well No. 4 in the  
Company or Operator Lease

NE 1/4 of Sec. 20, T. 21-S, R. 36-E, N. M. P. M.,  
Eunice Field, Lea County.

The dates of this work were as follows: 12-17-35

Notice of intention to do the work was [xxx] submitted on Form C-102 on 12-10-35 19  
and approval of the proposed plan was [was not] obtained. (Cross out incorrect words.)

DETAILED ACCOUNT OF WORK DONE AND RESULTS OBTAINED

THIS well was spudded in December 17, 1935.

Witnessed by \_\_\_\_\_  
Name Company Title

Subscribed and sworn to before me this \_\_\_\_\_

day of \_\_\_\_\_, 19 \_\_\_\_\_

Notary Public

My Commission expires \_\_\_\_\_

I hereby swear or affirm that the information given above is true and correct.

Name *A. G. Schuckler*

Position District Engineer

Representing Shell Petroleum Corp  
Company or Operator

Address Box P, Hobbs, N. M.

Remarks:

*A. J. [Signature]*  
Name

Title

THE UNIVERSITY OF CHICAGO  
DEPARTMENT OF CHEMISTRY

PHYSICAL CHEMISTRY

1. The rate of reaction between hydrogen and chlorine is first order in each reactant. The rate constant at 298 K is  $1.2 \times 10^{-7} \text{ s}^{-1}$ . Calculate the half-life of the reaction if the initial concentration of hydrogen is  $0.1 \text{ mol l}^{-1}$  and the initial concentration of chlorine is  $0.1 \text{ mol l}^{-1}$ .

2. The rate of reaction between hydrogen and chlorine is first order in each reactant. The rate constant at 298 K is  $1.2 \times 10^{-7} \text{ s}^{-1}$ . Calculate the half-life of the reaction if the initial concentration of hydrogen is  $0.1 \text{ mol l}^{-1}$  and the initial concentration of chlorine is  $0.1 \text{ mol l}^{-1}$ .

3. The rate of reaction between hydrogen and chlorine is first order in each reactant. The rate constant at 298 K is  $1.2 \times 10^{-7} \text{ s}^{-1}$ . Calculate the half-life of the reaction if the initial concentration of hydrogen is  $0.1 \text{ mol l}^{-1}$  and the initial concentration of chlorine is  $0.1 \text{ mol l}^{-1}$ .

4. The rate of reaction between hydrogen and chlorine is first order in each reactant. The rate constant at 298 K is  $1.2 \times 10^{-7} \text{ s}^{-1}$ . Calculate the half-life of the reaction if the initial concentration of hydrogen is  $0.1 \text{ mol l}^{-1}$  and the initial concentration of chlorine is  $0.1 \text{ mol l}^{-1}$ .

5. The rate of reaction between hydrogen and chlorine is first order in each reactant. The rate constant at 298 K is  $1.2 \times 10^{-7} \text{ s}^{-1}$ . Calculate the half-life of the reaction if the initial concentration of hydrogen is  $0.1 \text{ mol l}^{-1}$  and the initial concentration of chlorine is  $0.1 \text{ mol l}^{-1}$ .

6. The rate of reaction between hydrogen and chlorine is first order in each reactant. The rate constant at 298 K is  $1.2 \times 10^{-7} \text{ s}^{-1}$ . Calculate the half-life of the reaction if the initial concentration of hydrogen is  $0.1 \text{ mol l}^{-1}$  and the initial concentration of chlorine is  $0.1 \text{ mol l}^{-1}$ .

7. The rate of reaction between hydrogen and chlorine is first order in each reactant. The rate constant at 298 K is  $1.2 \times 10^{-7} \text{ s}^{-1}$ . Calculate the half-life of the reaction if the initial concentration of hydrogen is  $0.1 \text{ mol l}^{-1}$  and the initial concentration of chlorine is  $0.1 \text{ mol l}^{-1}$ .

8. The rate of reaction between hydrogen and chlorine is first order in each reactant. The rate constant at 298 K is  $1.2 \times 10^{-7} \text{ s}^{-1}$ . Calculate the half-life of the reaction if the initial concentration of hydrogen is  $0.1 \text{ mol l}^{-1}$  and the initial concentration of chlorine is  $0.1 \text{ mol l}^{-1}$ .

9. The rate of reaction between hydrogen and chlorine is first order in each reactant. The rate constant at 298 K is  $1.2 \times 10^{-7} \text{ s}^{-1}$ . Calculate the half-life of the reaction if the initial concentration of hydrogen is  $0.1 \text{ mol l}^{-1}$  and the initial concentration of chlorine is  $0.1 \text{ mol l}^{-1}$ .

10. The rate of reaction between hydrogen and chlorine is first order in each reactant. The rate constant at 298 K is  $1.2 \times 10^{-7} \text{ s}^{-1}$ . Calculate the half-life of the reaction if the initial concentration of hydrogen is  $0.1 \text{ mol l}^{-1}$  and the initial concentration of chlorine is  $0.1 \text{ mol l}^{-1}$ .