

(Nov. 1936)

Indian Agency.....

(SUBMIT IN TRIPLICATE)

U. S. Land Office 03 Las CrucesLease or permit No. 032099-B

Allottee.....

UNITED STATES
DEPARTMENT OF THE INTERIOR
GEOLOGICAL SURVEY

Lease No.

SUNDRY NOTICES AND REPORTS ON WELLS

NOTICE OF INTENTION TO DRILL.....	SUBSEQUENT REPORT OF WATER SHUT-OFF.....	<input checked="" type="checkbox"/>
NOTICE OF INTENTION TO CHANGE PLANS.....	SUBSEQUENT REPORT OF SHOOTING OR ACIDIZING.....	
NOTICE OF INTENTION TO TEST WATER SHUT-OFF.....	SUBSEQUENT REPORT OF ALTERING CASING.....	
NOTICE OF INTENTION TO RE-DRILL OR REPAIR WELL.....	SUBSEQUENT REPORT OF REDRILLING OR REPAIR.....	
NOTICE OF INTENTION TO SHOOT OR ACIDIZE.....	SUBSEQUENT REPORT OF ABANDONMENT.....	
NOTICE OF INTENTION TO PULL OR ALTER CASING.....	SUPPLEMENTARY WELL HISTORY.....	
NOTICE OF INTENTION TO ABANDON WELL.....		

(INDICATE ABOVE BY CHECK MARK NATURE OF REPORT, NOTICE, OR OTHER DATA)

Lockhart B-31 #4

Hobbs, N.M. November 28, 1938

Well No. 4 is located 330 ft. from ~~XXX~~ S line and 660 ft. from E line of sec. 31SE/4 Sec. 31
(1/4 Sec. and Sec. No.)21-S 36-E
(Twp.) (Range)N.M.P.M.
(Meridian)Eunice
(Field)Lea
(County or Subdivision)New Mexico
(State or Territory)

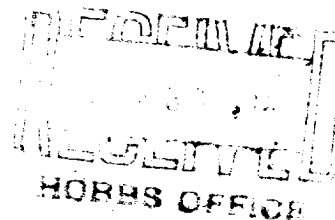
The elevation of the derrick floor above sea level is _____ ft.

DETAILS OF WORK

(State names of and expected depths to objective sands; show sizes, weights, and lengths of proposed casings; indicate mudding jobs, cementing points, and all other important proposed work)

After standing cemented 5 days, 5½" casing was tested with 1200# pressure for 30 minutes both before and after drilling plug. Casing tested OK.

FGP:



I understand that this plan of work must receive approval in writing by the Geological Survey before operations may be commenced.

Company Continental Oil CompanyAddress Box CC, Hobbs, N.M.By [Signature]Title District Supt.

THE UNIVERSITY OF CHICAGO
DEPARTMENT OF CHEMISTRY
RESEARCH REPORT NO. 10, 1954
BY J. H. GOLDSTEIN

THE EFFECT OF TEMPERATURE ON THE RATE OF
HYDROLYSIS OF ETHYL ACETATE

The rate of hydrolysis of ethyl acetate was measured at various temperatures. The reaction was carried out in a constant volume, constant pressure calorimeter. The rate of reaction was determined from the change in heat of reaction with time. The results are shown in the following table:

Temperature (°C)	Rate of Reaction (mole/liter-hr)
10	0.0012
20	0.0025
30	0.0050
40	0.0100
50	0.0200
60	0.0400
70	0.0800
80	0.1600
90	0.3200
100	0.6400

The rate of reaction increases with increasing temperature. The activation energy of the reaction was determined from the Arrhenius plot. The activation energy is 15.0 kcal/mole.

The rate of reaction is also affected by the concentration of the reactants. The rate of reaction increases with increasing concentration of ethyl acetate. The rate of reaction is also affected by the concentration of the catalyst. The rate of reaction increases with increasing concentration of the catalyst.

The rate of reaction is also affected by the presence of water. The rate of reaction increases with increasing concentration of water. The rate of reaction is also affected by the presence of other substances. The rate of reaction is also affected by the presence of other substances.

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