

Page 2  
New Mexico Oil Conservation Commission  
May 6, 1955

This request for enlargement of the present unit to a gas unit containing 156.65 acres is in accordance with the Commission rules and regulations and it is requested that this gas unit be granted.

Yours very truly,

P. J. DINGLADE, OPERATOR

By J. H. Moore  
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cj

Encl

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1. *Enthalpy of formation* of a compound is the heat change when one mole of the compound is formed from its elements in their standard states.

Example:

For the reaction:  $\text{C(s)} + \text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)}$   
 The enthalpy change is  $-393.5 \text{ kJ mol}^{-1}$ .  
 This is the enthalpy of formation of  $\text{CO}_2$  from its elements.

2. *Enthalpy of combustion* is the heat change when one mole of a substance is completely burnt in oxygen.

Example:



3.

4. *Enthalpy of solution* is the heat change when one mole of a substance is dissolved in a large volume of water.

Example:

5. *Enthalpy of hydration* is the heat change when one mole of an anhydrous salt is dissolved in a large volume of water.

Example:

6. *Enthalpy of neutralization* is the heat change when one mole of an acid reacts with one mole of an alkali to form one mole of water.