

## NEW MEXICO OIL CONSERVATION COMMISSION

SANTA FE, NEW MEXICO

## MISCELLANEOUS NOTICES

Submit this notice in triplicate to the Oil Conservation Commission or its proper agent before the work specified is to begin. A copy will be returned to the sender on which will be given the approval, with any modifications considered advisable, or the rejection by the Commission or agent, of the plan submitted. The plan as approved should be followed, and work should not begin until approval is obtained. See additional instructions in the Rules and Regulations of the Commission.

Indicate nature of notice by checking below:

|   |                                     |   |  |
|---|-------------------------------------|---|--|
| NOTICE OF INTENTION TO TEST CASING SHUT-OFF | <input checked="" type="checkbox"/> | NOTICE OF INTENTION TO SHOOT OR CHEMICALLY TREAT WELL |  |
| NOTICE OF INTENTION TO CHANGE PLANS         |                                     | NOTICE OF INTENTION TO PULL OR OTHERWISE ALTER CASING |  |
| NOTICE OF INTENTION TO REPAIR WELL          |                                     | NOTICE OF INTENTION TO PLUG WELL                      |  |
| NOTICE OF INTENTION TO DEEPEN WELL          |                                     |   |  |

Hobbs, New Mexico

Place

Oct. 1, 1948

Date

OIL CONSERVATION COMMISSION,  
Santa Fe, New Mexico.

Gentlemen:

Following is a notice of intention to do certain work as described below at the \_\_\_\_\_  
Skelly Oil Company H. O. Sims "D" Well No. 2 in SE/4 NW/4  
 Company or Operator Lease  
 of Sec. 3, T. 23 S, R. 37 E, N. M. P. M., Drinkard Field.  
Lea County.

## FULL DETAILS OF PROPOSED PLAN OF WORK

FOLLOW INSTRUCTIONS IN THE RULES AND REGULATIONS OF THE COMMISSION

Drilled to T.D. 2790', ran and cemented string of 8-5/8" OD casing at 2790' with 1000 sacks of cement. Halliburton process used. Will let set 60 hours and on Oct. 3, 1948, at approximately 3:00 P.M., will drill plug and test casing shut-off.

Approved \_\_\_\_\_  
 except as follows:

OIL CONSERVATION COMMISSION,

By

Title

Skelly Oil Company

Company or Operator

By

Position Dist. Supt.

Send communications regarding well to

Name

Skelly Oil Company

Address Drawer "D"

Hobbs, New Mexico

THE UNIVERSITY OF CHICAGO

DEPARTMENT OF CHEMISTRY

PHYSICAL CHEMISTRY

1. The following data were obtained from a study of the reaction of hydrogen peroxide with various metal ions in aqueous solution at 25°C. The rate of reaction was measured by the disappearance of hydrogen peroxide over a period of 10 minutes.

2. The rate of reaction was found to be first order with respect to the concentration of hydrogen peroxide and first order with respect to the concentration of the metal ion. The rate constant,  $k$ , was determined for each metal ion and is given in the table below. The units of  $k$  are  $\text{min}^{-1}$ .

3. The following data were obtained from a study of the reaction of hydrogen peroxide with various metal ions in aqueous solution at 25°C. The rate of reaction was measured by the disappearance of hydrogen peroxide over a period of 10 minutes.

4. The rate of reaction was found to be first order with respect to the concentration of hydrogen peroxide and first order with respect to the concentration of the metal ion. The rate constant,  $k$ , was determined for each metal ion and is given in the table below. The units of  $k$  are  $\text{min}^{-1}$ .

5. The following data were obtained from a study of the reaction of hydrogen peroxide with various metal ions in aqueous solution at 25°C. The rate of reaction was measured by the disappearance of hydrogen peroxide over a period of 10 minutes.

6. The rate of reaction was found to be first order with respect to the concentration of hydrogen peroxide and first order with respect to the concentration of the metal ion. The rate constant,  $k$ , was determined for each metal ion and is given in the table below. The units of  $k$  are  $\text{min}^{-1}$ .

7. The following data were obtained from a study of the reaction of hydrogen peroxide with various metal ions in aqueous solution at 25°C. The rate of reaction was measured by the disappearance of hydrogen peroxide over a period of 10 minutes.

8. The rate of reaction was found to be first order with respect to the concentration of hydrogen peroxide and first order with respect to the concentration of the metal ion. The rate constant,  $k$ , was determined for each metal ion and is given in the table below. The units of  $k$  are  $\text{min}^{-1}$ .

9. The following data were obtained from a study of the reaction of hydrogen peroxide with various metal ions in aqueous solution at 25°C. The rate of reaction was measured by the disappearance of hydrogen peroxide over a period of 10 minutes.

10. The rate of reaction was found to be first order with respect to the concentration of hydrogen peroxide and first order with respect to the concentration of the metal ion. The rate constant,  $k$ , was determined for each metal ion and is given in the table below. The units of  $k$  are  $\text{min}^{-1}$ .